

## Investigación

# Study of the Influence of the Polyacetylene Main Chain Configuration and the Internal Stacking of Pendant Groups over the Optical and Photophysical Properties of *trans* and *cis*-Poly(1-pyrenylacetylene)s by UV-Vis and Fluorescence Spectroscopy

Ernesto Rivera\*

Instituto de Investigaciones en Materiales, Universidad Nacional Autónoma de México, Circuito Exterior, Ciudad Universitaria C.P. 04510 México D.F. México. Tel.: (52-55) 56 22 47 33; fax: (55-25) 56 16 12 01. e-mail riverage@zinalco.iimatercu.unam.mx

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**Abstract.** The influence of the polyacetylene backbone configuration and the internal stacking of pendant groups over the optical and photophysical properties of *trans*-poly(1-ethynylpyrene) (*trans*-PEP) and *cis*-poly(1-ethynylpyrene) (*cis*-PEP) was studied by UV-vis and fluorescence spectroscopy. From the absorption spectra of both polymers, it was found that *trans*-PEP possesses a higher degree of conjugation than its *cis* analogue. Besides, intramolecular interactions occur between the pendant pyrene units (associated pyrenes) present in each polymer, giving rise to a new emission. These interactions are stronger in *cis*-PEP due to a coiling of the main polyacetylene chain. Excitation spectra jointly with fluorescence decay profiles show that these interactions occur in the ground state (static excimer).

**Key words:** Intramolecular interactions, polyarylacetylenes, pyrene, UV-vis, fluorescence.

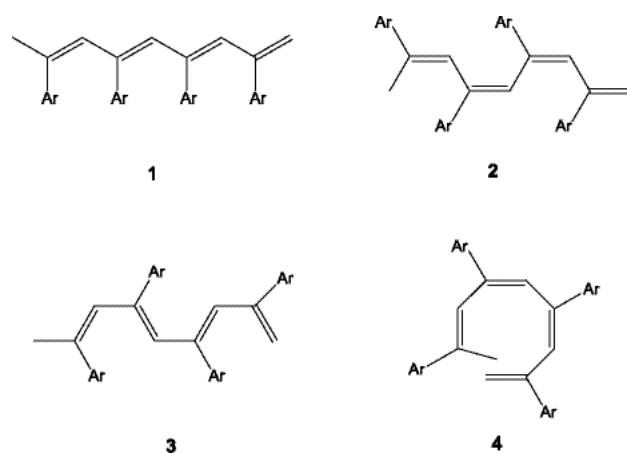
**Resumen.** Se estudió la influencia de la configuración de la cadena principal de poliacetileno y el apilamiento interno de los grupos pendientes en las propiedades ópticas y fotofísicas del *trans*-poli(1-ethinilpireno) (*trans*-PEP) y el *cis*-poli(1-ethinilpireno) (*cis*-PEP) por espectroscopía UV-vis y de fluorescencia. A partir de los espectros de absorción de ambos polímeros, se encontró que el *trans*-PEP posee un grado de conjugación más alto que su análogo *cis*. Además, en cada polímero tienen lugar interacciones intramoleculares entre las unidades de pireno pendientes (pirenos asociados), dando lugar a una nueva emisión. Estas interacciones son más fuertes en *cis*-PEP debido al pliegue de la cadena principal de poliacetileno en sí misma. Los espectros de excitación conjuntamente con los perfiles de desactivación de fluorescencia mostraron que estas interacciones ocurren en el estado fundamental (excímero estático).

**Palabras clave:** Interacciones intramoleculares, poliarilacetilenos, pireno, UV-vis, fluorescencia.

## Introduction

$\pi$ -Conjugated polymers have been regarded in the last 30 years as promising materials for the development of electronic devices such as light emitting diodes, photovoltaic cells and non linear optical systems [1-5]. The opto-electronic properties vary significantly depending on the degree of extended conjugation between the consecutive repeat units. In particular, polyacetylene is the most conjugated polymer and shows metallic conductivity upon doping. However its insolubility, instability as well as its improcessability limit its practical applications as a functional material [6]. Introduction of alkyl or aryl substituents into polyacetylenes, can improve remarkably its stability. However, this increase in stability is often accompanied by a decrease in the degree of conjugation due to torsions of the polymer backbone, skewing it from planarity due to steric interactions between pendant substituents [7-10].

Polyarylacetylenes have been widely studied by many research groups, specially polyphenylacetylene [11]. In fact, there are four possible geometries for polyarylacetylenes (Fig. 1): *trans*-transoidal (1), *cis*-transoidal (2), *trans*-cisoidal (3) and *cis*-cisoidal (4). For polyphenylacetylene, all these geometries can be easily differentiated by IR and  $^1\text{H-NMR}$  spectroscopies [12-13].



**Fig. 1.** Possible geometries for polyarylacetylenes: *trans*-transoidal (1), *cis*-transoidal (2), *trans*-cisoidal (3), *cis*-cisoidal (4).

Arylacetylenes bearing non-polar groups are usually polymerised in the presence of  $\text{WCl}_6$  and disubstituted acetylenes with  $\text{TaCl}_5$  [14]. Masuda and co-workers [15] carried out the polymerisation of several arylacetylenes bearing bulky condensed aromatic rings obtaining polymers with molecular weights up to 140 000 g/mol depending on the structure of the mono-

mer. Surprisingly, these polymers showed to be highly conjugated in despite of enhanced steric effects. The position of the acetylene bond in the aromatic ring influences significantly the reactivity of the monomer towards polymerisation due to steric and electronic effects [15]. Nevertheless, arylacetylenes bearing polar groups such as -CN, -OR, -SR, NR and -N=N- cannot be polymerised with  $\text{WCl}_6$  [16] but they can be polymerised using other catalytic systems [16-19].

In the present work, we study the influence of the polyacetylene backbone geometry and the internal stacking of pendant substituents on polyarylacetylenes bearing condensed aromatic groups. We selected poly(1-pyrenylacetylene) or poly(1-ethynylpyrene) as model because pyrene has been widely used as probe in many fluorescence studies on a huge variety of labeled polymers and biopolymers [20-29]. Pyrene has a long single lifetime and readily form excimers. The most important features of the photophysics of pyrene in various environments were discussed in detail by F.M. Winnik in a recent review [30].

In a previous study [31-32], we polymerised 1-ethynylpyrene (EP) or 1-pyrenylacetylene and other monomer derivatives trying different reaction conditions. Poly(1-ethynylpyrene)s (PEP) obtained using  $\text{WCl}_6$  as catalyst, sometimes in the presence of a co-catalyst, gave molecular weights from 24 000 to 470 000 g/mol and polydispersities from 2.9 to 11. *Trans*-poly(1-ethynylpyrene)s (*trans*-PEP) (Fig. 2) obtained by this method possess *trans-transoidal* geometry were dark purple polymers, totally soluble in *o*-dichlorobenzene, very soluble in THF and partially soluble in toluene and  $\text{CHCl}_3$  [31]. On the other hand, EP was also polymerised in the presence of  $\text{Rh}_2\text{nbd}_2\text{Cl}_2$  to give *cis*-poly(1-ethynylpyrene). The obtained polymer was dark brown and practically insoluble in all common organic solvents [15]. It is very well known that this behaviour is characteristic for *cis-cisoidal* polyarylacetylenes. Tabata and co-workers studied a similar polymer, poly(1-naphthylacetylene), obtained with the same catalyst by Resonance Raman spectroscopy and it was found to possess *cis-cisoidal* geometry [33].

Since *cis-cisoidal* polymers are insoluble and difficult to characterized by the common methods, we carried out the polymerisation of EP trying a new catalytic system (1-Me-Indenyl)( $\text{PPh}_3$ ) $\text{Ni-C}\equiv\text{C-Ph}$  and methylaluminoxane (MAO), (named here NiC/MAO) [32] developed by Zargarian and co-workers [34]. *Cis*-poly(1-ethynylpyrene)s (*cis*-PEP) obtained with this catalyst were soluble in various organic solvents and possessed *cis-transoidal* geometry (Figure 2). Molecular weights  $M_w$  from 2 200 to 24 000 g/mol and polydispersities about 2 were obtained with this catalytic system. *Cis*-PEP obtained with NiC/MAO were dark brown polymers, totally soluble in *o*-dichlorobenzene, fairly soluble in THF and poorly soluble in  $\text{CHCl}_3$ . Since the optical properties of polyarylacetylenes depend on a big measure of the geometry of the polyacetylene main chain, we decided to carry out a comparative study between *trans*-PEP synthesized with  $\text{WCl}_6$  and *cis*-PEP obtained with NiC/MAO, bearing identical molecular weights ( $M_w = 24 000$  g/mol).

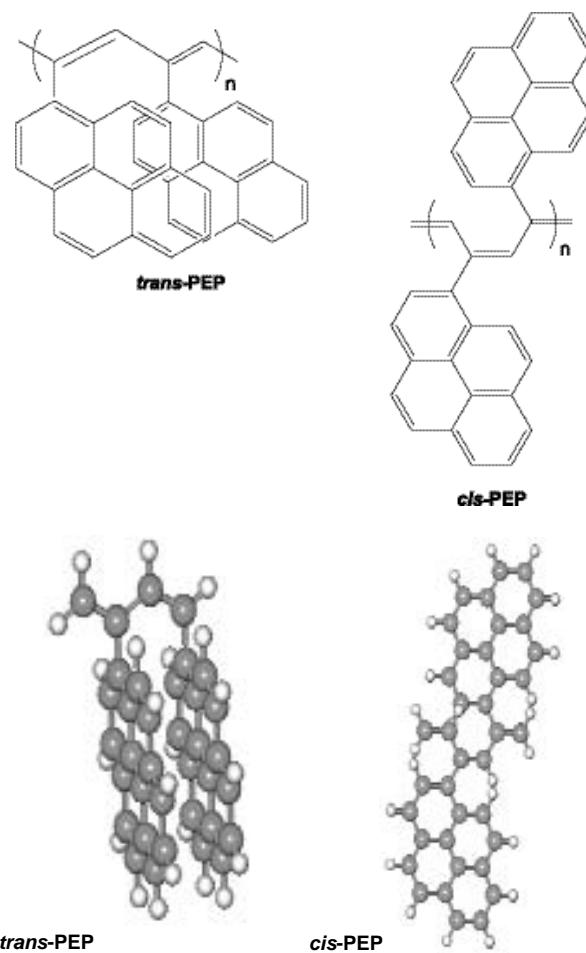


Fig. 2. Structures of *trans*-PEP and *cis*-PEP.

## Experimental Part

The synthesis and characterization of *trans*-poly(1-ethynylpyrene) (*trans*-PEP) [31] and *cis*-poly(1-ethynylpyrene) (*cis*-PEP) [32] was achieved according to our previously published procedures. Polymers having identical molecular weights ( $M_w = 24 000$  g/mol) were selected for this work. The chemical structures of *trans*-PEP and *cis*-PEP are shown in Fig. 2.

For UV-vis and fluorescence spectroscopies, tetrahydrofuran (THF) was purchased from Aldrich (spectrophotometric grade). Prior to use, the solvent was checked for spurious emission in the region of interest and found to be satisfactory. Pyrene was purified by crystallization from hexanes prior to be used. The absorption spectra in solution and in the solid state (casted film from THF solution) were recorded on a Varian Cary 1 Bio UV/vis spectrophotometer (model 8452A) using 1 cm quartz cells and solute concentrations of  $1-3 \times 10^{-5}$  M for the polymers. It has been verified that the Beer-Lambert law applies for the concentrations used. Fluorescence spectra corrected for the emission detection were recorded on a Spex Fluorolog-2 spectrophotometer with an F2T11 special configuration. Each solution was excited near the absorption wave-

length maximum using a 1 cm quartz cell. Concentrations  $5.4 \times 10^{-3}$  and  $1.45 \times 10^{-6}$  M were used for pyrene solutions, whereas for the polymers a concentration about  $1.3 \times 10^{-6}$  M was used giving absorbances less than 0.1 in order to avoid any inner filter effect.

Fluorescence lifetimes were measured on a multiplexed time-correlated single photon counting fluorometer (Edinburgh Instruments, model 299T). The instrument incorporates an all-metal coaxial hydrogen flash-lamp. Reconvolution analysis was performed by fitting over all the fluorescence decay including the rising edge. The kinetic interpretation of the goodness-of-fit was assessed using plots of weighted residuals, reduced  $\chi^2$  value, and Durbin-Watson (DW) parameters. The lifetime measurements were performed in argon-saturated solutions of the substrates at room temperature.

## Results and Discussion

### Absorption spectra of pyrene derivatives

The absorption spectrum of pyrene in THF diluted solution (not shown) exhibits three sets of absorption bands around 336, 273 and 241 nm corresponding to the  $S_2$ ,  $S_3$  and  $S_4$  excited states. However, the absorption band due to  $S_1$  is so weak that it can be seen only at high concentrations and appears around  $\lambda = 372$  nm. According to the Kasha's rule, fluorescence comes exclusively from the  $S_1$  state and higher excited states have to decay to  $S_1$  before light emission takes place [35]. Thus, it is more practical to excite at the wavelength corresponding to the  $S_2$  band, which is more visible. Pyrene polymers usually exhibit this band around  $\lambda = 336$  nm. For all the molecules studied in this work, this band appears at  $\lambda = 346$  nm, 10 nm red-shifted compared to that of pyrene itself. Therefore, in the present study we recorded the fluorescence spectra of DEP, *trans*-PEP and *cis*-PEP exciting at this wavelength in order to obtain maximum emission.

### Characterization, thermal and optical properties of *trans*-PEP and *cis*-PEP

*Trans* and *cis*-PEP were previously characterized by IR, solid state  $^1\text{H}$  and  $^{13}\text{C}$ -NMR and Resonance Raman spectroscopies [31-33]. Thermal properties of these polymers were determined by TGA from 0 to 1000 °C and DSC from -40 to 400 °C [32]. According to the absorption spectra of *trans*-PEP and *cis*-PEP in THF (Figure 3), *trans*-PEP exhibit a peak at 336 nm which is attributed to the pyrene groups in the non-associated state. Another peak at 346 nm has its origin in intramolecular interactions between adjacent pyrene units (associated pyrenes) along the polymer backbone. Finally, the broad band observed at 580 nm is caused by the polyacetylene chain, which indicates that the effective electronic conjugation is longer for this polymer. This was confirmed because the absorption spectrum of a *trans*-PEP derivative with DP = 2, di(1-ethynylpyrene) (DEP), did not show any absorption band at this wavelength (Figure 3). ZINDO/S calculations confir-

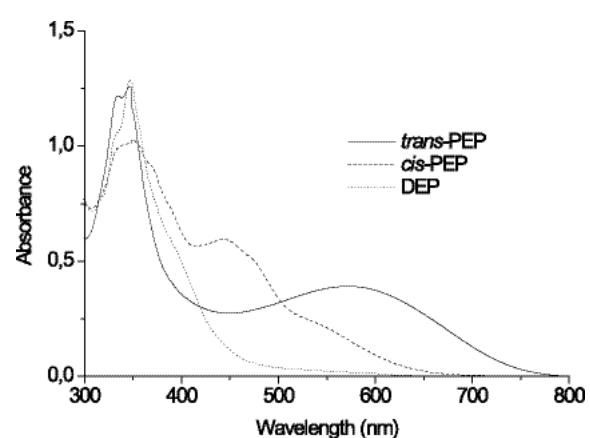


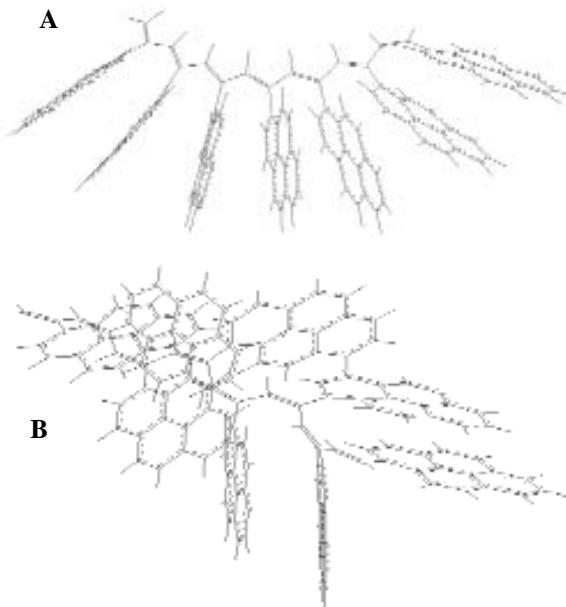
Fig. 3. Absorption spectra for *trans*-PEP (24 mg/L), *cis*-PEP (24 mg/L) and DEP (16 mg/L) in THF solution.

med that 10 to 12 coplanar repetitive units can reproduce the energy of this band [36].

On the other hand, the absorption spectrum of *cis*-PEP is very different from that observed for *trans*-PEP. Indeed, a new band appears at 453 nm, which clearly indicates that another kind of intramolecular interactions take place in *cis*-PEP. Moreover, the band observed at 580 nm in the absorption spectrum of *trans*-PEP is not observed for *cis*-PEP. Instead, a shoulder at 537 nm, which cannot be attributed to the absorption spectrum of associated pyrene units. This shoulder is due to the polyacetylene main chain and is an indication that the effective degree of conjugation is shorter for *cis*-PEP. Previously, molecular mechanics calculations (MM2) performed on short segments of *trans*-PEP and *cis*-PEP (Fig. 4) [32] confirmed that for the *trans*-PEP the polymer backbone is well aligned with pyrene pendant units perpendicular to the polyacetylene main chain plane. Pyrene groups adopted an almost parallel orientation giving regioregular intramolecular stacking. By contrast, in *cis*-PEP (Figure 4) steric interactions between pyrene groups and neighbour hydrogen groups force the main chain to twist out of plane thereby reducing the effective conjugation length. This gives rise to a partial and non-regioregular stacking of the pyrene units. Thus, in *cis*-PEP pendant pyrene groups interact in a non-parallel way. Distances between pyrene units varies from 3.1 Å (near the backbone) to 4.1 Å (far away from it) in the *trans*-PEP and in *cis*-PEP were estimated to be from 4.4 to 5.3 Å in a selected almost parallel pyrene-pyrene complex. Nevertheless, it can be shorter or longer along the *cis*-polyacetylene backbone for other pyrene pairs. It is worth to point out that in *cis*-PEP, pyrene units are not well aligned bearing a random orientation.

### Brief aspects of pyrene excimers

An excimer was defined by Birks as a dimer which is associated in an electronic excited state and which is dissociative in its ground state [35]. The formation of a pyrene excimer

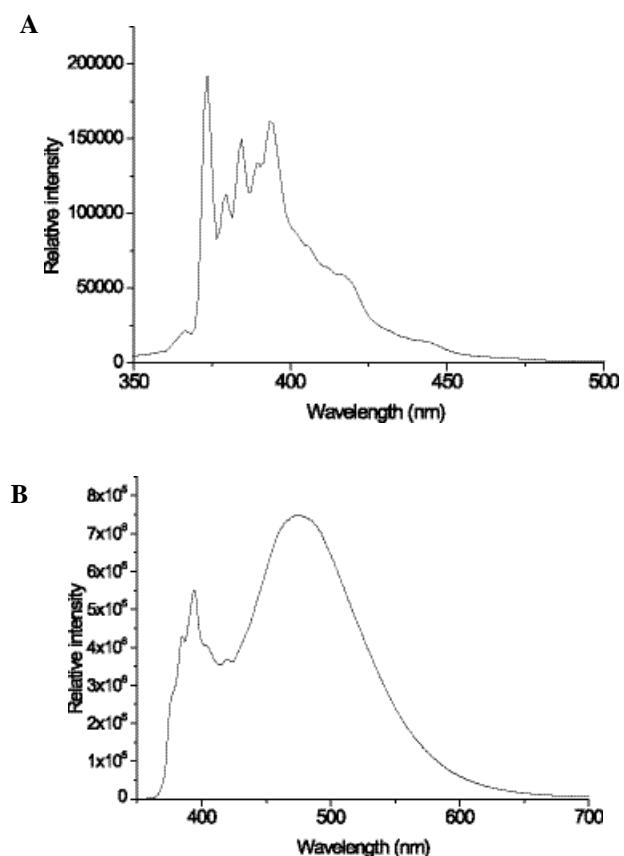


**Fig. 4.** Modelling structure for short segments of *trans*-PEP and *cis*-PEP obtained by molecular mechanics calculations (MM2).

requires encounter by diffusion of an electronically excited pyrene with a second pyrene in its ground state. According to this definition the two pyrenes must be sufficiently far apart when light is absorbed, so that the excitation is localized only on one of them. This excited pyrene often referred as “locally excited”, named also “excited pyrene in the non-associated state” gives rise to a “monomer emission” or “non-associated pyrene emission”. Since both pyrene units have to move in order to interact, the emission resulting from this process is known as “excimer” or “dynamic excimer”. For intermolecular interactions, high concentration ensures the encounter of two pyrene molecules by diffusion since they are closer, whereas in diluted solutions pyrene units are considerably far apart to interact. Figure 5 shows fluorescence spectra of pyrene in THF at low and high concentration. In diluted solution we can observe a well structured emission band of excited pyrene in the non-associated state at 385-464 nm with a cut off about 470 nm. By contrast, in concentrated solution, besides this band an excimer emission band at 476 nm is also observed, due to intermolecular interactions between pyrene units.

There are also instances where a like-excimer emission is observed, when two pyrenes are close to each other forming a complex before light is absorbed. Because this emission strongly resembles a pyrene excimer emission and pyrene units did not have to move to interact, many people name these excited species “static excimers” or “excited complexes” [30].

Mainly, there are two ways to distinguish between dynamic and static excimers: analysing excitation spectra and fluorescence life time decay profiles. For dynamic excimers, if we record excitation spectra at the fluorescence wavelength ( $\lambda_F$ )



**Fig. 5.** Fluorescence spectra of pyrene in THF: a) at low concentration (0.3 mg/L, excitation at  $\lambda = 336$  nm). b) at high concentration (1.1 g/L, excitation at  $\lambda = 336$  nm).

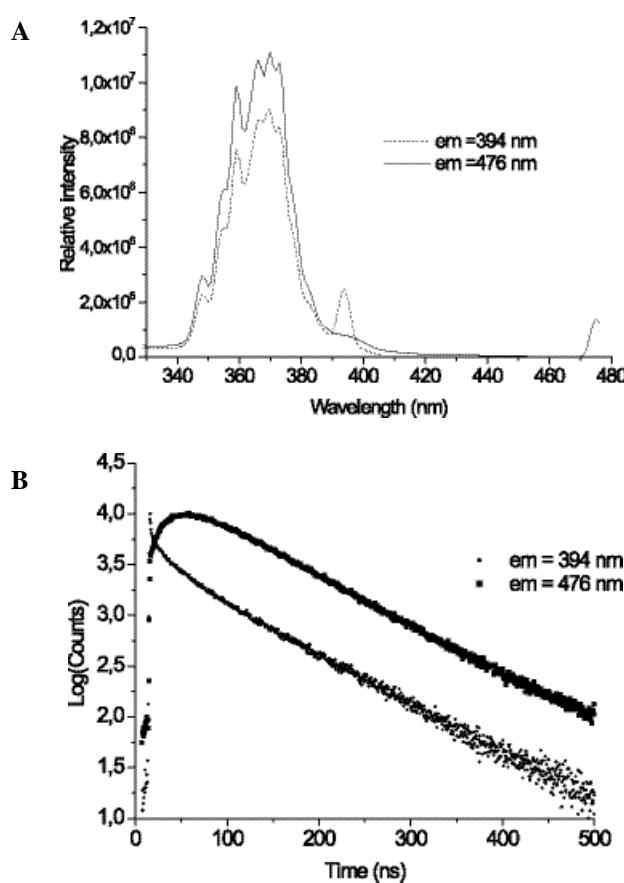
of excited non-associated pyrene and at  $\lambda_F$  of the excimer emission, both excitation spectra have to be identical (Fig. 6). On the other hand, if we record the fluorescence life time decay profile at  $\lambda_F$  of excited non-associated pyrene and at  $\lambda_F$  of the excimer emission, in the latter we can observe a rising component for the excimer formation followed by a negative slope showing fluorescence decay and in the former only decay. Regarding static excimers, if we record the fluorescence life time decay profile at  $\lambda_F$  of excited non-associated pyrene and at  $\lambda_F$  of the like-excimer emission, in both cases only a negative slope or fluorescence decay is observed (Fig. 6).

The classic Birks two-state mechanism for excimer formation and decay is expressed in two biexponential equations: (1) for the monomer emission and (2) for the excimer formation.

$$I_M = A_1 e^{-t/\tau_1} + A_2 e^{-t/\tau_2} \quad (1)$$

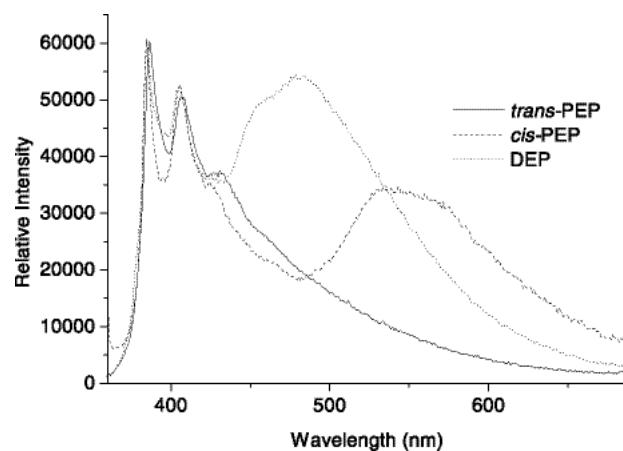
$$I_E = A_3 e^{-t/\tau_3} - A_4 e^{-t/\tau_4} \quad (2)$$

According to these equations the monomer or the excited non-associated pyrene and the excimer are coupled species  $\tau_1 = \tau_3$ ,  $\tau_2 = \tau_4$  and  $A_4/A_3 = 1$ . For a dynamic excimer this model applies and as we can see  $-A_4$  is a negative coefficient. However, for static excimers deconvolution analysis gives a positive  $A_4$  value [30].



**Fig. 6.** Pyrene at high concentration (1.1 g/L): a) Excitation spectra (emission at  $\lambda_F = 387$  nm and  $\lambda_F = 476$  nm). b) Fluorescence life time decay profiles (emission at  $\lambda_F = 387$  nm and  $\lambda_F = 476$  nm).

*Steady state fluorescence spectra of trans-PEP and cis-PEP.* The fluorescence spectrum of *trans*-PEP (Fig. 7) in THF solution shows a band in the 360-465 nm region, due to non-associated pyrene units. Apparently, it does not show any excimer emission but it shows an important emission intensity near 500 nm, a wavelength where pyrene itself does not emit. According to the photophysical behaviour of pyrene, this provides evidence of the formation of a excited-complex between pyrene units. By contrast, the fluorescence spectrum of *cis*-PEP in the same solvent shows two distinct bands (Fig. 7). The first band, with vibronic peaks, was very similar to that observed for *trans*-PEP and could be attributed to non-associated pyrene units. The broad band around 550 nm results from the formation of intramolecular interactions between adjacent pyrene units along the molecular frame. Indeed, similar spectral properties observed in the fluorescence spectrum of *cis*-PEP were previously reported by us for di(1-ethynyl-pyrene) (DEP) (Fig. 7). The spectral position of this band is very different from that observed for DEP ( $\lambda_F \approx 480$  nm) because of the different conformation adopted by the main chain in *cis*-PEP. Indeed, in *cis*-PEP intramolecular interactions formed by a coiling of the main chain take place giving

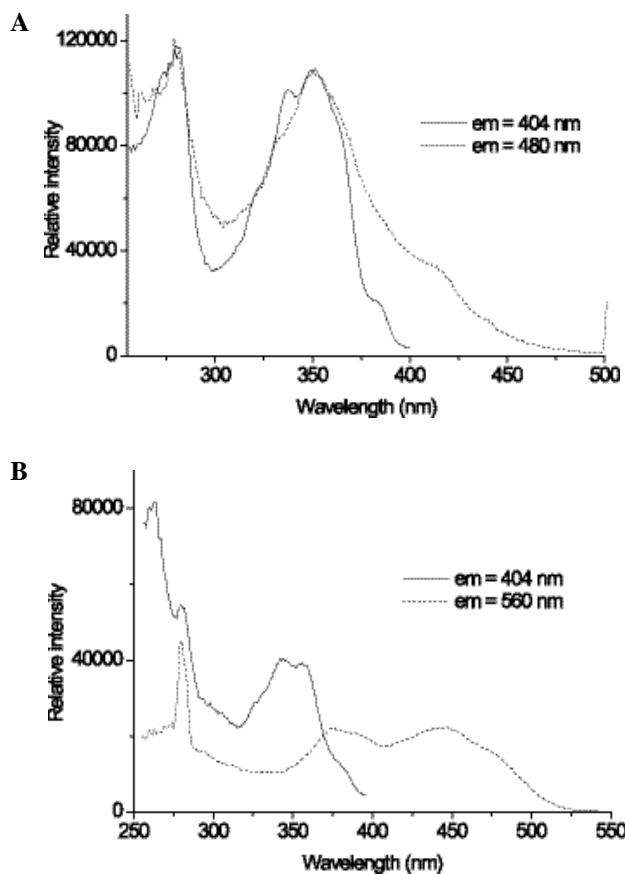


**Fig. 7.** Fluorescence spectra of *trans*-PEP (1.5 mg/L, excitation  $\lambda = 346$  nm), *cis*-PEP (2.5 mg/L, excitation at  $\lambda = 346$  nm) and DEP (2 mg/L, excitation  $\lambda = 346$  nm) in THF solution.

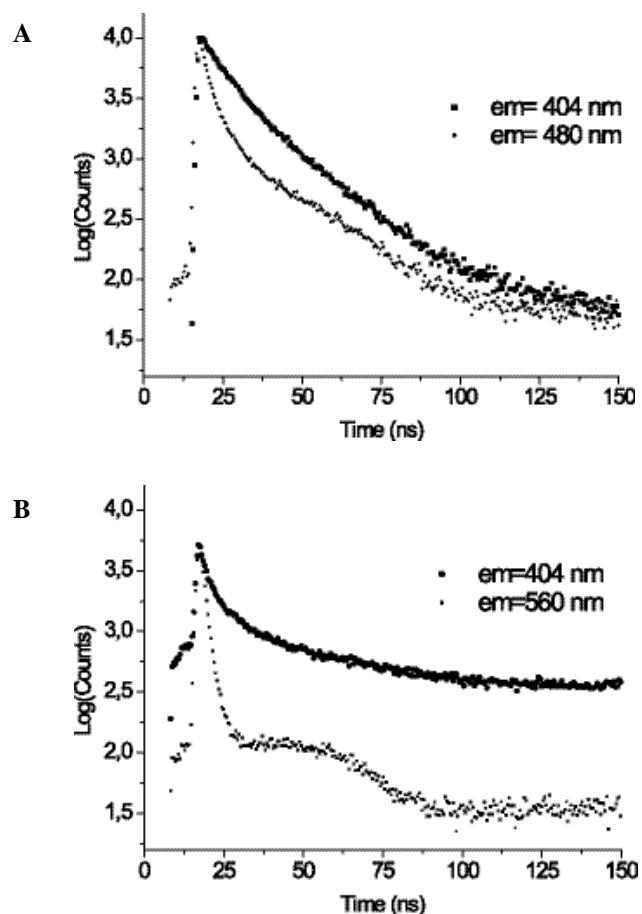
rise to shorter distances for several pyrene units, and consequently to a large red-shift of this emission band.

*Excitation spectra of trans-PEP and cis-PEP.* In order to differentiate between the formation of a dynamic or static excimer, excitation spectra of *trans*-PEP were recorded at  $\lambda_F = 404$  nm and at  $\lambda_F = 480$  nm and showed to be quite different. Excitation spectra were measured for *cis*-PEP at  $\lambda_F = 404$  nm and  $\lambda_F = 560$  nm as well (Figure 8). The excitation spectrum measured at  $\lambda_F = 560$  nm is very different from that measured at  $\lambda_F = 404$  nm. This is against the rule of the formation of a dynamic excimer, which should exhibit an excitation spectrum identical to that of the non-associated pyrene units. The band centered at 385 nm and the large band around 446 nm correspond to the shoulder at 365 nm and the band at 453 nm observed in the absorption spectrum of *cis*-PEP (Fig. 3). However, the shoulder observed at 537 nm in the absorption spectrum of *cis*-PEP is not observed in the excitation spectrum of this compound. This clearly shows that this band was not caused by pyrene groups (associated or non-associated) and involved only the polyacetylene main chain as mentioned above.

*Lifetime fluorescence decay profiles of trans-PEP and cis-PEP.* The fluorescence decay profiles of *trans*-PEP and *cis*-PEP, recorded at the non-associated pyrene  $\lambda_F$  and at the excimer fluorescence wavelength (Figure 9) provided interesting results of the kinetic analysis and are summarized in Table 1. The best fits obtained for these decays are complex and involve a triexponential expression. This is due to the presence of various species (arrangements) along the polymer backbones. The fluorescence decay recorded at the longer wavelength of each polymer does not show any rising component, which should be observed for an emission originating from a dynamic excimer [30]. Moreover, no negative component was calculated from the decay profile analysis, which is against the formation of a dynamic excimer in *trans*-PEP and *cis*-PEP.



**Fig. 8.** Excitation spectra of *trans*-PEP (1.5 mg/L, emission at  $\lambda_F = 404$  nm and  $\lambda_F = 480$  nm) and *cis*-PEP (2.5 mg/L, emission at  $\lambda_F = 404$  nm and  $\lambda_F = 560$  nm) in THF.



**Fig. 9.** Lifetime decay fluorescence profiles for *trans*-PEP (1.5 mg/L, emission at  $\lambda_F = 404$  nm and  $\lambda_F = 480$  nm) and *cis*-PEP (2.5 mg/L, emission at  $\lambda_F = 404$  nm and  $\lambda_F = 560$  nm) in THF.

Thus, we can conclude that we have static excimers in both polymers and interactions between adjacent pyrene groups take place in the ground state.

The average lifetime measured for *trans*-PEP and *cis*-PEP (Table 1) at the longer wavelengths are significantly shorter than that measured about  $\lambda_F = 404$  nm. This is a proof that a new static complex (static excimer) between the pendant pyrene units present in the polymer appear at longer wavelengths. Surprisingly, the fluorescence lifetime of non-associated pyrene units is unusually low. Indeed, the fluorescence lifetime of the pyrene chromophore is normally quite high, reaching 450 ns in non-polar solvents such as cyclohexane and decreasing drastically depending on the polarity of the solvent [37]. Recently, Borgig *et al* reported that pyrene chromophores grafted to polyethylene showed similar short fluorescence lifetimes [38]. For instance, 1-pyrenylbutyl methacrylate (PyBMA) grafted to polyethylene, which have a fluorescence spectrum (385 nm, 406 nm) very close to that of our polymers, exhibits a lifetime of 13.3 ns.

Therefore, the average fluorescence lifetime obtained about  $\lambda_F = 404$  nm for *cis*-PEP is much shorter than that of *trans*-PEP measured at the same wavelength. This is a proof

that pyrene units in the *cis*-PEP are in a different arrangement, being closer to each other by a coiling of the main chain (Figure 4). Similarly, the average lifetime calculated at  $\lambda_F = 560$  nm is shorter than that of *trans*-PEP at  $\lambda_F = 480$ , suggesting that this complex is stronger than that formed with pyrene units in the *trans*-polymer. All these facts are in agreement with the spectral band positions of both complexes, indicating that the complex present in the *cis*-polymer is more stabilized.

## Conclusion

Optical properties of *trans*-PEP and *cis*-PEP in THF solution were investigated by UV-vis and fluorescence spectroscopy. *Trans*-PEP possesses an effective degree of conjugation, which is higher than that of *cis*-PEP. Both polymers show emissions originating from excited non-associated pyrene units and static excimers. From excitation spectra and fluorescence decay profiles measured at the maximum of both fluorescence bands, it is evident that ground state interactions (static excimers) are responsible for the resulting long wavelength

**Table 1.** Fluorescence decay parameters of *trans*-PEP and *cis*-PEP in THF at room temperature.

Polymer	$\lambda_{\text{exc}}$ (nm)	$\lambda_{\text{F}}$ (nm)	$\tau_{\text{F}}$ (ns)	$\mathbf{B}^{\text{b}}$	$f^{\text{d}}$	$\chi^2$
<i>trans</i> -PEP	346	404	$\tau_1^{\text{a}} = 3.0$	$B_1 = 0.24$	0.06	1.40
			$\tau_2^{\text{a}} = 12.9$	$B_2 = 0.73$	0.81	
<i>cis</i> -PEP	346	480	$\tau_3^{\text{a}} = 59.8$	$B_3 = 0.03$	0.13	1.08
			$\langle \tau_{\text{F}} \rangle^{\text{c}} = 18.3$			
<i>trans</i> -PEP	346	404	$\tau_1^{\text{a}} = 1.5$	$B_1 = 0.71$	0.31	1.28
			$\tau_2^{\text{a}} = 6.2$	$B_2 = 0.22$	0.38	
<i>cis</i> -PEP	346	560	$\tau_3^{\text{a}} = 17.8$	$B_3 = 0.07$	0.31	1.18
			$\langle \tau_{\text{F}} \rangle^{\text{c}} = 8.4$			
<i>trans</i> -PEP	346	404	$\tau_1^{\text{a}} = 0.2$	$B_1 = 0.95$	0.34	
			$\tau_2^{\text{a}} = 2.1$	$B_2 = 0.03$	0.15	
<i>cis</i> -PEP	346	560	$\tau_3^{\text{a}} = 13.5$	$B_3 = 0.02$	0.51	
			$\langle \tau_{\text{F}} \rangle^{\text{c}} = 7.2$			
<i>trans</i> -PEP	346	560	$\tau_1^{\text{a}} = 0.2$	$B_1 = 0.78$	0.32	
			$\tau_2^{\text{a}} = 1.4$	$B_2 = 0.21$	0.59	
<i>cis</i> -PEP	346	560	$\tau_3^{\text{a}} = 6.7$	$B_3 = 0.01$	0.09	
			$\langle \tau_{\text{F}} \rangle^{\text{c}} = 1.5$			

a) Obtained from the deconvolution fit:  $A + B_1 \exp(-t/\tau_1) + B_2 \exp(-t/\tau_2) + B_3 \exp(-t/\tau_3)$ . b) Normalized preexponential factor. c) Average lifetime calculated from the expression:  $\langle \tau_{\text{F}} \rangle = \sum_i B_i \tau_i^2 / \sum_i B_i \tau_i$ . d) Fractional contribution to the total fluorescence intensity defined as:  $f_i(\lambda) = B_i \tau_i / \sum_i B_i \tau_i$ .

emissions. It is worth pointing out that the long-wavelength fluorescence band of *cis*-PEP is more stabilized, suggesting that pyrene units are closer to each other in this polymer and interact in a non-parallel way.

The fluorescence decay profiles of both polymers are complex and are best fitted by a triexponential expression. Since the average fluorescence lifetime of *cis*-PEP is much shorter than that of *trans*-PEP, stronger interactions between pyrene units (correlated to shorter intermolecular distances) in the *cis*-PEP are responsible for this behaviour.

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