

# Water adsorption on rutile titanium dioxide (110): Theoretical study of the effect of surface oxygen vacancies and water flux in the steady state case

F. Bouzidi<sup>a,d,\*</sup>, M. Tadjine<sup>a</sup>, A. Berbri<sup>a</sup> and A. Bouhekkab<sup>b,c</sup>

<sup>a</sup>*Department of Physics, Faculty of Exact Sciences and Informatics, University Hassiba Ben Bouali of Chlef, P. B. 78 C, National Road N° 19, Ouled Fares 02180, Chlef, Algeria.*

<sup>b</sup>*Department of Materials Sciences, Faculty of Sciences and Technology, Tissemsilt University, Algeria.*

<sup>c</sup>*Thin-Film Physics Laboratory and Materials for Electronics, Oran 1 Ahmed Ben Bella University, P.B. 1524, El M'naouar 31000, Oran, Algeria.*

<sup>d</sup>*Laboratory of Mechanics and Energy, University Hassiba Ben Bouali of Chlef, Hay Salem, National Road N° 19, 02000, Chlef, Algeria.*

\*e-mail: bouzidifatima94@gmail.com

Received 16 May 2022; accepted 4 November 2022

The aim purpose of the present work is highlighting the impact of surface oxygen vacancies and H<sub>2</sub>O flux on the behavior of water adsorption at the rutile titanium dioxide (110). Therefore, a theoretical model, based on molecular and dissociation mechanisms at different surface atomic sites, was formulated in a system of partial differential coupled equations. The proposed model used to study, in an atomic scale, this complex phenomenon of adsorption governed by several factors including surface vacancies defects and water flux. The findings of the solution of the system of equations in the steady state case, presented in this paper, strongly indicated that the rate coverage of surface oxygen vacancies has an important role in the dissociation of H<sub>2</sub>O as well as the flux which is a key factor in the behavior of water adsorption on the rutile TiO<sub>2</sub> (110) and the rate coverage of OH groups.

**Keywords:** Water adsorption; rutile titanium dioxide (110); oxygen vacancies; hydroxyls groups; H<sub>2</sub>O flux; steady state.

DOI: <https://doi.org/10.31349/RevMexFis.69.031004>

## 1. Introduction

Among a lot of important materials, titanium dioxide (TiO<sub>2</sub>) takes a particular attention in the recent years in several industrial applications especially in the elimination of organic contamination, sensor, splitting of water, self-cleaning and photocatalytic [1–5]. Furthermore, TiO<sub>2</sub> is abundance [3], inexpensive, nontoxic [4], inert and chemical stability [5, 6]. Rutile TiO<sub>2</sub> (110) surface is the lowest surface energy, thus is very important in studying, especially in catalysis [7]. Furthermore, TiO<sub>2</sub> has a high refractive index ranges between 2.4–2.6. Unfortunately, one of the drawbacks of TiO<sub>2</sub> is the wide band gap of 3–3.2 eV, reacts only in the (UV) light under (380 nm), that is available only around 5% in solar energy [8–11]. Many researchers treated this deficiency by generating electronic states inside the forbidden gap via inducing defects (oxygen vacancy, Ti interstitial, dislocation,...) and doping with impurities (metal transition, noble metal,...) in the structures of TiO<sub>2</sub>. This leads to extend its response in the visible light and near infrared (NIR) region [12–17]. One of the most important applications of TiO<sub>2</sub> surface is water splitting. The H<sub>2</sub>O can be adsorbed in two different mechanisms either molecular or dissociative onto rutile TiO<sub>2</sub> (110) surface [18, 19]. But this is an intriguing phenomenon, because the adsorption process is affected by many factors (phase components, temperature, vacancies, water coverage...) [20]. Several theoretical methods (Langmuir, Freundlich, Langmuir-

Hinshelwood...) are widely used to study the kinetics of adsorptions [21–25]. These phenomena have been studied by using various ab initio quantum mechanical approximations highlighting on the important factors that affect on the adsorption and the dissociation of H<sub>2</sub>O such as (DFT and DFT-PW and HF...) calculations and molecular dynamics ([21] and references therein) which have provide numerous groundbreaking insight into the water -TiO<sub>2</sub> interface. However, it is not possible to get adsorption information taking in consideration all the surface sites at the same time. Experimentally, the infrared spectroscopy, especially in Situ ATR-FTIR, is one of the most important and powerful technique used to investigate the water molecules adsorption at solid surfaces [26]. But in experimental, it is not clear and easy to get details at the atomic scale level of adsorption therefore; the theory becomes a useful tool. In light of the above information, we are interested in this context, to study the effect of oxygen vacancies concentration at rutile TiO<sub>2</sub> (110) surface and water flux on the behavior of H<sub>2</sub>O molecules adsorption using a system of nonlinear differential equations, carried out taking into account the different cases of adsorption and all surface available sites. This indicates a better understanding of structural and dynamical properties changes of the behavior of H<sub>2</sub>O on rutile TiO<sub>2</sub> (110) surface. Thereby, these factors greatly enhancing chemical/catalytic reactivity (among other properties) as well as the splitting of water and self-cleaning. At the end, a summary will be given together with

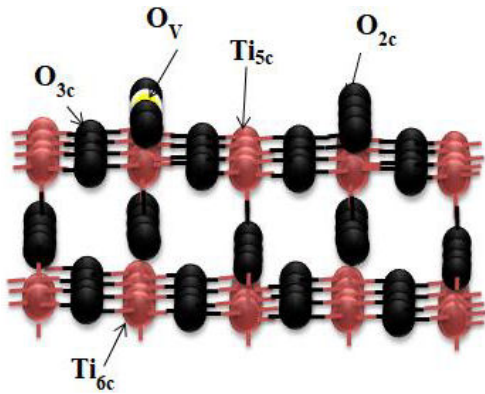


FIGURE 1. Structure of rutile  $\text{TiO}_2$  (110) surface in the z-direction, black balls represent O atoms of the bridge-bonded  $\text{O}_{2c}$  and three coordinated  $\text{O}_{3c}$ , red balls are fivefold coordinated surface Ti ( $\text{Ti}_{5c}$ ) and six fold coordinated Ti atoms ( $\text{Ti}_{6c}$ ), white ball is single oxygen vacancies ( $\text{O}_V$ ).

a brief outlook towards challenges and prospects for future theoretical studies [27,28].

## 2. The surface structure of rutile $\text{TiO}_2$ (110)

Naturally,  $\text{TiO}_2$  exists in three different crystallography forms rutile, anatase and brookite. The structure of each form is described as a chain of octahedron (one atom of Ti surrounded by six oxygen atoms) as shown in Fig. 1. Among the three structures anatase (101) and rutile (110) are the most stable faces and more reactive especially in photocatalytic applications [7].

The different surface sites illustrated in this figure play a crucial role in the adsorption and the behavior of water molecules. It is still not clear to distinguish which kind of site is primordial in controlling the adsorption phenomenon. In the following part, we give the possible mechanisms of adsorption used to formulate the proposed theoretical model.

## 3. Kinetics model description of $\text{H}_2\text{O}$ - rutile $\text{TiO}_2$ (110) surface interactions

A kinetics model of  $\text{H}_2\text{O}$  on rutile  $\text{TiO}_2$  (110) surface, in bellow figures, was developed to describe the behavior of  $\text{H}_2\text{O}$  adsorption, assumed to adsorb in reversible states; different forms model, two types of features found on rutile  $\text{TiO}_2$  (1x1) (110) surface as indicated in Fig. 2 and reaction (1).  $\text{H}_2\text{O}$  molecules adsorb in molecular form on defect-free sites ( $\text{Ti}_{5c}$ ) at low temperature around 160 K [29] is due a Ti - $\text{H}_2\text{O}$  bond length (2.16 - 2.29) Å [30], with adsorption energy around (0.5 - 0.7 eV) [1,28,33–35]. When annealing the surface below room temperature, only the  $\text{H}_2\text{O}$  molecules are desorbed from surface with energy ranges between 0.7 eV to 0.8 eV depending on the water rate coverage [36]. In the other hand, The  $\text{H}_2\text{O}$  has the possibility to dissociate on ( $\text{Ti}_{5c}$ ) sites, however, the second proton  $\text{H}^+$  transfer to the neighbouring bridging oxygen ( $\text{O}_{2c}$ ) and forms two OH

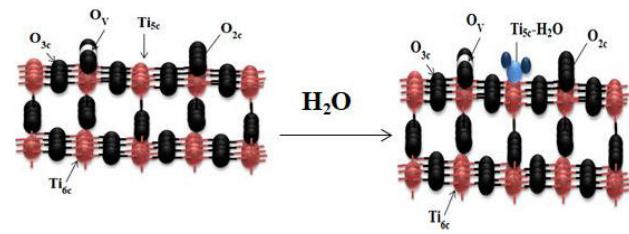


FIGURE 2. The adsorption of molecular water on free defect surface, dark (light) blue balls hydrogen (oxygen of water) atoms,  $K_1$ ,  $K_2$  are the rates constants of adsorption and desorption respectively.

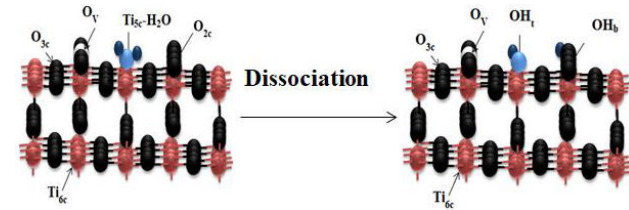
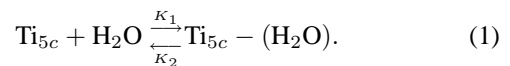
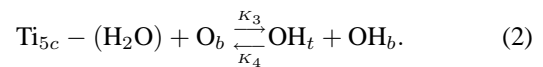


FIGURE 3. The dissociative adsorption of water on  $\text{Ti}_{5c}$  sites, where  $K_3$  and  $K_4$  are the rates constants of dissociation and recombination reaction, respectively.

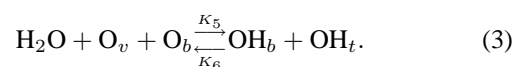
hydroxyl groups  $\text{OH}_t$  and  $\text{OH}_b$  respectively, which was discussed in details in Ref. [3] with dissociation energy toward 0.36 eV and 0.41 eV, when annealing the surface to room temperature under reaction (2) [31–33] as shown in Fig. 3. This competition between molecular and dissociated adsorption is mostly governed by the basic strength of the bridging O atom receiving the H.



A missing of an oxygen atom, by removing the oxygen bridging ( $\text{O}_{2c}$ ) from the surface using electron bombardment or other methods, leads to the creation of a surface oxygen vacancy ( $\text{O}_V$ ) (none thermally). This was discussed in more details in Ref. [2].



With its concentration around 15% ML (Mono Layer) (1 ML is defined with respect to the coverage of  $\text{Ti}_{5c}$  sites of  $5.2 \cdot 10^{14} \text{ cm}^{-2}$ ) from the surface of  $\text{TiO}_2$  (110) [34,35]. The water favoured dissociate in the oxygen defect at below temperature (130-200 K) and forms two OH hydroxyl groups in the  $\text{O}_V$  and the  $\text{H}^+$  proton of water transfer nearby bridging oxygen and creates the other hydroxyl for every vacancy as given in Fig. 4 [36–39], the recombination of the two OH occurs at temperature above 490 K [3,18]. This reaction can be written as Eq. (3) indicates.



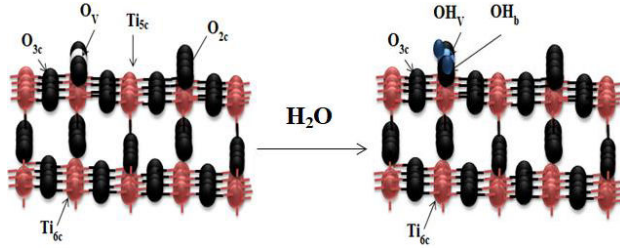


FIGURE 4. The dissociative adsorption of water on the surface oxygen vacancy, where the  $K_5$  and  $K_6$  are the dissociation and recombination rates constants, respectively

#### 4. Model and method

A mathematical description of these equations (1)-(3) for the adsorption of water in different cases onto rutile  $\text{TiO}_2$  (110) surface can be derived using an adaptation of the Langmuir adsorption model. This model is the most commonly applied of single component liquid-solid adsorption. Based on the kinetics principle, Langmuir isotherm model assuming that only monolayer adsorption exists; no impurities ( $\text{CO}_2$ ,  $\text{O}_2$ ...) are considered at the surface and to avoid interactions between water molecules the incident  $\text{H}_2\text{O}$  flux density can be supposed very weak and more addition the adsorbent surface is uniform with the same adsorption probability [22,24]. Therefore, the adsorption rate of each chemical reaction can be presented as following Eq. (4):

$$\frac{d\theta}{dt} = \Phi P_{\text{ads}} - K\theta, \quad (4)$$

where  $\Phi$ ,  $\theta$ ,  $P_{\text{ads}}$ ,  $K$  are the flux, the surface sites coverage, the probability of the molecule will find adsorption sites and the desorption rate constant, respectively. The probability of adsorption is given by the following Eq. (5).

$$P_{\text{ads}} = K \left( \frac{n_a(1 - \theta_a)}{n_s} \right), \quad (5)$$

where  $K$ ,  $n_a$ ,  $n_s$ ,  $\theta_a$  the adsorption rate constant ( $\text{s}^{-1}$ ), the concentration of adsorption sites, the concentration of all atoms in the surface and the coverage of the adsorption sites

respectively. The rate constant  $K_i$  can be expressed with an Arrhenius Eq. (6),

$$K_i = K_0 \exp \left( \frac{-E_a}{RT} \right), \quad (6)$$

where  $K_0$ ,  $E_a$ ,  $R$  and  $T$  are the attempt frequency ( $\text{s}^{-1}$ ), the activation energy (Kj/mole,  $1\text{eV} = 96.482\text{Kj}$ ), universal gas constant ( $8.314\text{J/mole.K}$ ) and the temperature K.

The different reactions (1) to (3) using Eq. (4) are organized below,

$$\text{Ti}_{5c} + \text{H}_2\text{O} \xrightleftharpoons[\frac{K_2}{K_1}]{} \text{Ti}_{5c} - (\text{H}_2\text{O})$$

$$r_1 : \Phi K_1 \theta_{\text{Ti}} - K_2 \theta_{\text{H}_2\text{O}}, \quad (7)$$

$$\text{O}_b + \text{Ti}_{5c} - (\text{H}_2\text{O}) \xrightleftharpoons[\frac{K_4}{K_3}]{} 2\text{OH}$$

$$r_2 : K_3 \theta_{\text{H}_2\text{O}} \theta_b - K_4 \theta_{\text{OH}}^2, \quad (8)$$

$$\text{O}_v + \text{O}_b + (\text{H}_2\text{O}) \xrightleftharpoons[\frac{K_6}{K_5}]{} 2\text{OH}$$

$$r_3 : \Phi K_5 \theta_v \theta_b - K_6 \theta_{\text{OH}}^2, \quad (9)$$

where  $\theta_{\text{Ti}}$ ,  $\theta_{\text{OH}}$ ,  $\theta_{\text{H}_2\text{O}}$ ,  $\theta_b$  and  $\theta_v$  are the coverage of titanium, hydroxyl groups, water, bridging oxygen and oxygen vacancies on the surface respectively. The adsorption rate is proportional to the  $\text{H}_2\text{O}$  flux arrived on surface and the coverage of active sites ( $\text{Ti}_{5c}$ ,  $\text{O}_v$  and  $\text{O}_{2c}$ ), desorption rate is proportional to the number of association and desorption adsorbed molecules, under such assumptions; in the adsorption process we have [39].

$$\frac{d\theta_v}{dt} = -r_3 = -\Phi K_5 \theta_v \theta_b + K_6 \theta_{\text{OH}}^2,$$

$$\frac{d\theta_b}{dt} = -r_3 - r_2 = -\Phi K_5 \theta_v \theta_b + K_4 \theta_{\text{OH}}^2 - K_3 \theta_{\text{H}_2\text{O}} \theta_b + K_6 \theta_{\text{OH}}^2,$$

$$\frac{d\theta_{\text{OH}}}{dt} = +r_3 + r_2 = \Phi K_5 \theta_v \theta_b - K_4 \theta_{\text{OH}}^2 + K_3 \theta_{\text{H}_2\text{O}} \theta_b - K_6 \theta_{\text{OH}}^2,$$

TABLE I. Kinetic parameters used in the mathematical models of  $\text{H}_2\text{O}$  reactions on rutile  $\text{TiO}_2$  (110), the rate constant ( $K_i$ ) extracted from the Arrhenius Eq. (6).

Step,(i)	$K_0[\text{s}^{-1}]$	$E_a$ [eV]	$T[\text{K}]$	$K_0$ [ $\text{s}^{-1}$ ]	Reference
1. $\text{H}_2\text{O}$ adsorption	$10^{13}$	0.5 – 0.7	150 – 160	0.0018	[1,3]
2. $\text{H}_2\text{O}$ desorption	$10^{12}$	0.73 – 0.8	200 – 275	$4.10^{-7}$	
3. $\text{H}_2\text{O}$ dissociation on free defect	$10^{12}$	0.36, 0.44, 0.97	80 – 140	0.0015	[29,30]
4.OH Association	$10^{12}$	0.355	110 – 130	$5.43.10^{-5}$	
5. $\text{H}_2\text{O}$ dissociation on $\text{O}_v$	$10^{13}$	0.93 – 1.5	300, 187	0.0024	[41–47], [51,54]
6.OH recombination	$10^8$	0.12 – 0.18	450 – 500	$10^{-4}$	

$$\begin{aligned}\frac{d\theta_{Ti}}{dt} &= -r_1 = -\Phi K_1 \theta_{Ti} + K_2 \theta_{H_2O} \\ \frac{d\theta_{H_2O}}{dt} &= r_1 - r_2 = \Phi K_1 \theta_{Ti} - K_2 \theta_{H_2O} \\ &\quad - K_3 \theta_{H_2O} \theta_b + K_4 \theta_{OH}^2\end{aligned}\quad (10)$$

The sum of all the different coverage at the rutile (110) surface are  $\sum \theta_i = 1$  and the variation equal

$$\sum \frac{d\theta_i}{dt} = 0 \quad (11)$$

We suppose the variation of the coverage's remains the same in duration of the reaction, setting these rates to zero in the steady state gives the system (12) as indicated bellow. The stability of the set of equations is very important because any bifurcation parameter can dramatically change it and to avoid such behavior we took the rates constants  $K_i$  have similar thermal stability.

$$\begin{aligned}-\Phi K_5 \theta_v \theta_b + K_6 \theta_{OH}^2 &= 0 \\ -\Phi K_5 \theta_v \theta_b + K_4 \theta_{OH}^2 - K_3 \theta_{H_2O} \theta_b + K_6 \theta_{OH}^2 &= 0 \\ \Phi K_5 \theta_v \theta_b - K_4 \theta_{OH}^2 + K_3 \theta_{H_2O} \theta_b - K_6 \theta_{OH}^2 &= 0 \\ -\Phi K_1 \theta_{Ti} + K_2 \theta_{H_2O} &= 0 \\ \Phi K_1 \theta_{Ti} - K_2 \theta_{H_2O} - K_3 \theta_{H_2O} \theta_b + K_4 \theta_{OH}^2 &= 0\end{aligned}\quad (12)$$

## 5. Results and discussion

We use a numerical method to solve this system Eq. (12) and the corresponding solution of coverages is given by the following equations:

$$\begin{aligned}\theta_{OH} &= 0.16(-750\Phi\theta_v + 2.24 \\ &\quad \times \sqrt{1875\Phi\theta_v - 2237\Phi\theta_v^2 + 96210\Phi^2\theta_v^2}),\end{aligned}\quad (13)$$

$$\theta_{Ti} = 0.2\theta_v, \quad (14)$$

$$\theta_{H_2O} = 8.7\Phi\theta_v, \quad (15)$$

$$\theta_{O_v} = \frac{5.10^{-4}(10^3\Phi\theta_{O_v} - 2.10^3\Phi\theta_{O_v}^2 + 2.10^5\Phi^2\theta_{O_v}^2 - 670.8\Phi\theta_{O_v}\sqrt{\Phi\theta_{O_v}(1.8.10^3 - 2.2.10^3\theta_{O_v} + 9.6.10^4\Phi\theta_{O_v})})}{\Phi\theta_{O_v}}. \quad (16)$$

It is clear that the solutions of each surface site strongly depends on the flux of water and the surface oxygen vacancies ( $O_V$ ). As indicated in Fig. 5 (a), OH hydroxyls increases and reaches a maximum for  $O_V = 0.04$  then it decreases to the minimum for a higher value of  $O_V$ . The same behavior for OH hydroxyls is illustrated in Fig. 5 (b) when this flux changes from 0 to 1. This graph indicates that the coverage of OH hydroxyls increases at low flux values and reaches a maximum around 0.7 (for a flux value around 0.4) afterward it decreases which means that OH hydroxyls production is slightly reduced.

To illustrate theses variations of OH hydroxyls on the surface versus  $H_2O$  flux and  $O_V$ , a three dimension (3D) is shown in Fig. 6. The coverage of OH hydroxyls grows with increasing the  $O_V$  and reaches to a maximum around 0.7 ML at  $O_V$  equal 0.05 ML. At high coverage of  $O_V$  up to 10 % the production of OH hydroxyls decays and gets a small value almost zero. In

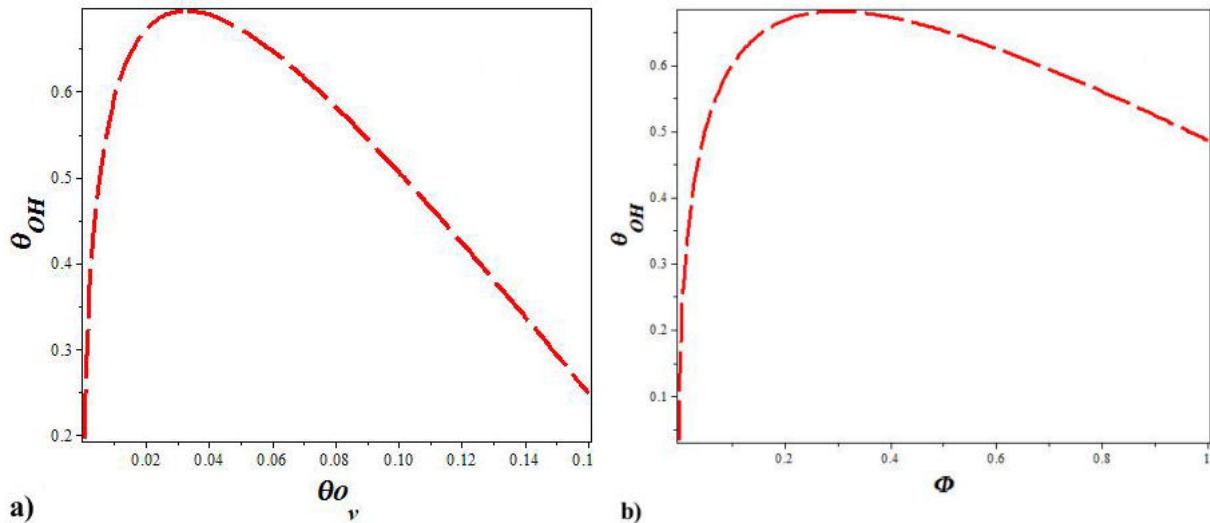


FIGURE 5. a) The effect of oxygen vacancies on the production OH hydroxyls at  $H_2O$  flux equal 0.4. b) The effect of  $H_2O$  flux on the production OH hydroxyls at  $O_v$  around 0.05 in the steady state solution.

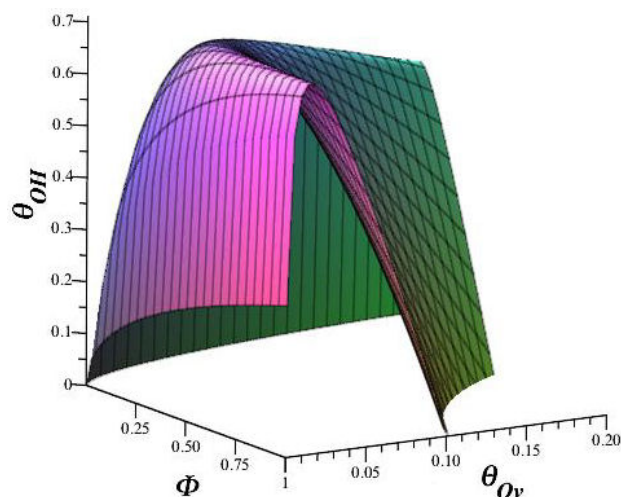
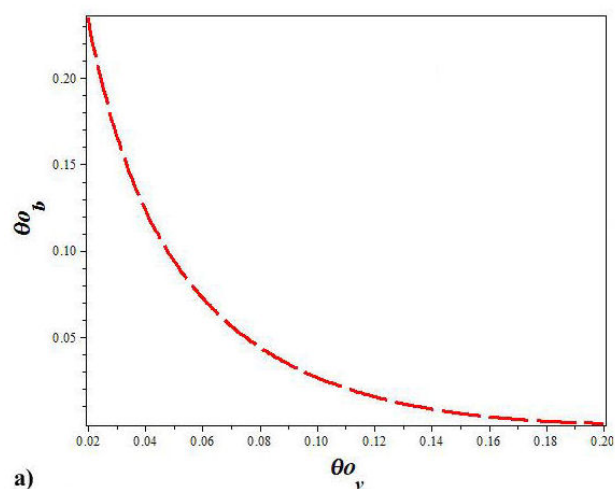
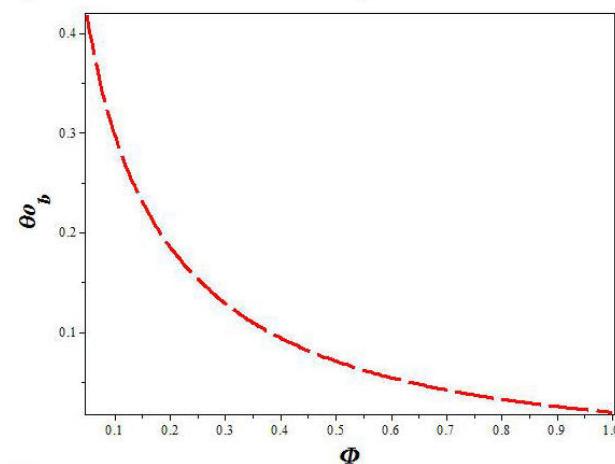


FIGURE 6. Kinetic curves for the OH hydroxyls at rutile  $\text{TiO}_2$  (110) surface for various  $\text{H}_2\text{O}$  flux and concentration of  $\text{O}_v$  on the surface.



a)



b)

FIGURE 7. The coverage of bridging oxygen change versus: a) the coverage of  $\text{O}_v$  with the  $\text{H}_2\text{O}$  flux remains 0.4 ML, b)  $\text{H}_2\text{O}$  flux at the coverage of  $\text{O}_v$  around 0.05 ML on rutile  $\text{TiO}_2$  (110) surface.

other hand, also the coverage of OH hydroxyls depends on the arriving  $\text{H}_2\text{O}$  molecules; when the  $\text{H}_2\text{O}$  flux arrived at surface increases, the production of OH hydroxyls increases and reaches a maximum at 0.7 ML for  $\text{H}_2\text{O}$  flux around 0.4 ML, despite, for  $\text{H}_2\text{O}$  flux takes value more than 0.4 ML and filling all the surface leading to no production of OH hydroxyls on the surface which is occupied only with  $\text{H}_2\text{O}$  molecules.

From Eq. (16), the curve of the coverage of  $\text{O}_b$  for different concentrations of  $\text{O}_v$  and  $\text{H}_2\text{O}$  flux arrived on surface is shown in Fig. 7.

Figure 7 (a) illustrates the variation of  $\theta_{\text{O}_b}$  as a function of  $\theta_{\text{O}_v}$  for the  $\text{H}_2\text{O}$  flux taken 0.4 ML. The graph clearly shows that  $\theta_{\text{O}_b}$  decreases when the concentration of  $\text{O}_v$  grows. Besides, in Fig. 7(b) the coverage of  $\text{O}_b$  decays exponential-like with increasing the  $\text{H}_2\text{O}$  flux for the concentration of  $\theta_{\text{O}_v}$  equals 0.05 ML.

Figure 8 is plotted using Eq. (16); it is interesting to note some common relations; the coverage of  $\text{O}_b$  decreases exponential-like with increasing the coverage of  $\text{O}_v$  and

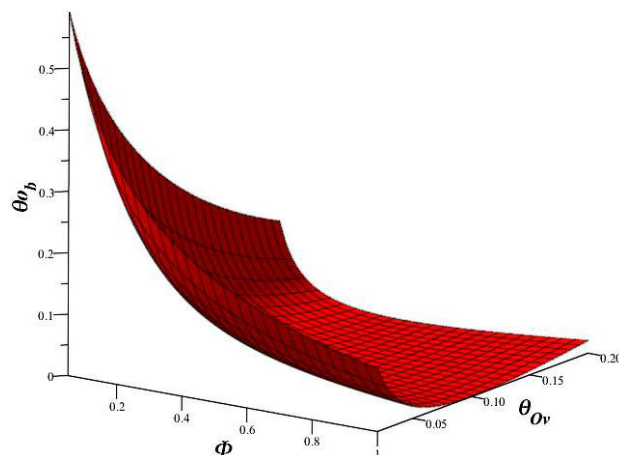


FIGURE 8. The coverage of  $\text{O}_b$  as a function of  $\text{H}_2\text{O}$  flux and  $\text{O}_v$  defects on rutile  $\text{TiO}_2$  (110) surface.

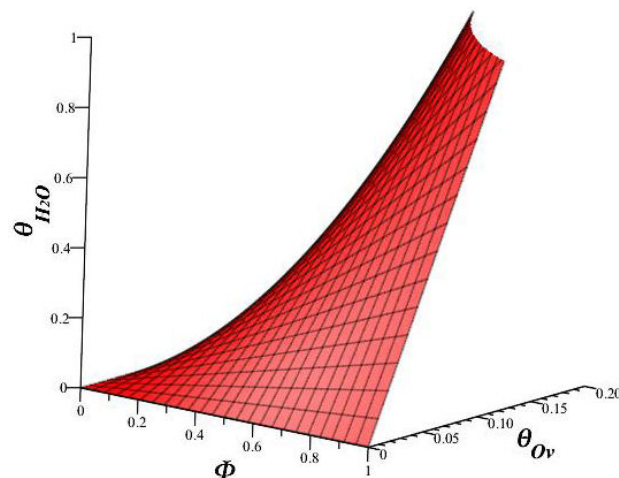


FIGURE 9. The coverage of  $\text{H}_2\text{O}$  on surface versus  $\text{H}_2\text{O}$  flux and the coverage of oxygen vacancies  $\text{O}_v$ .



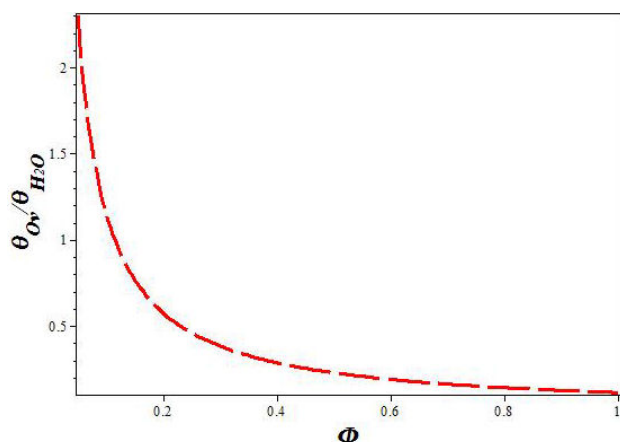


FIGURE 10. The ratio of  $(\theta_{O_v} / \theta_{H_2O})$  on rutile  $TiO_2$  (110) surface versus  $H_2O$  flux.

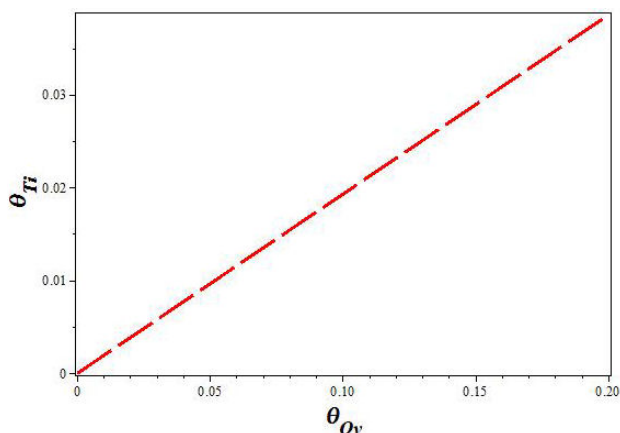


FIGURE 11. The coverage of  $Ti_{5c}$  on rutile  $TiO_2$  (110) surface as a function of the concentration of oxygen vacancies ( $O_v$ ).

tends to a small value for large value of  $O_v$ , also at high value for the  $H_2O$  flux, the coverage of  $O_b$  decreases.

This figure demonstrates that the  $O_b$  is a key factor in the dissociation and the adsorption of  $H_2O$  molecules.

The Eq. (15), variation of water coverage versus the rate of oxygen vacancies and the  $H_2O$  flux, indicating, the coverage of  $H_2O$  takes small values almost plateau when the  $H_2O$  flux and the concentration of  $O_v$  on surface is very small, afterward, the coverage of  $H_2O$  increases steadily with  $H_2O$  flux and the concentration of  $O_v$ , as illustrated in Fig. 9.

Figure 10 shows the ratio of  $(\theta_{O_v} / \theta_{H_2O})$ , from Eq. (15), versus the  $H_2O$  flux; indicating that the fraction of  $(\theta_{O_v} / \theta_{H_2O})$  decreases with increasing the  $H_2O$  flux. This can be explained that increasing the coverage of  $H_2O$  at the surface is caused by the filling of oxygen vacancies sites.

From these results, the interactions of  $H_2O$  with rutile  $TiO_2$  (110) surface were heavily affected by the  $O_v$  and  $H_2O$  flux. This was expected, since the  $H_2O$  flux arrived on rutile  $TiO_2$  (110) surface,  $H_2O$  molecules take up to  $Ti_{5c}$  sites (which is positively charged) rendering its able to bind  $H_2O$  molecules via electrostatic interactions. Many authors

[52, 54] have studied the effect of  $H_2O$  flux and the coverage of  $O_v$  on the behavior of  $H_2O$  molecules on rutile  $TiO_2$  (110) surface. First, at low values of  $H_2O$  flux, is expected the mobility of  $H_2O$  molecules to be high which is helpful for the  $H_2O$  molecules to diffuse on the surface on direction [001] which dissociate at oxygen vacancies firstly to heal the vacancies and defect-free sites as well as a strong attractive of H proton to near bridging oxygen due to transfer the H proton and formed another OH hydroxyls, because dissociation at bridging oxygen vacancies is more favourable than  $H_2O$  adsorption on the  $Ti_{5c}$  sites [55], as the  $H_2O$  flux increases the OH hydroxyls continue to increase till reach the maximum at 0.7ML for 0.4 of  $H_2O$  flux, this is in good accordance with the prediction by Lindan and Zhang [56]. At a high  $H_2O$  flux, water may be adsorbed molecularly with a small fraction adsorbed dissociative on  $O_v$  because (i) at least one or two  $H_2O$  molecules uptake on  $Ti_{5c}$  sites and form a chains of  $H_2O$  molecules (ii) when the amount of  $H_2O$  molecules are large, hinders the diffusion of  $H_2O$  molecules and remains in molecular forms, therefore,  $H_2O$  molecular adsorption is more favourable at high coverage [30]. As it is seen, the coverage of  $O_v$  affects on the adsorption of water, however, the amount of OH hydroxyls formed on rutile  $TiO_2$  (110) surface increases at low concentration of vacancies because the  $H_2O$  favours the dissociation at oxygen vacancies [54] reach to the maximum around 0.7 ML because the repulsive interaction between the functional groups, afterward, when the coverage of  $O_v$  increases beyond to 0.05 ML, the OH groups coverage decreases, because at high concentration of  $O_v$  affects on the structure properties, which lead to (2x1) reconstruction surface [40], additionally, and we know the  $O_v$ 's created from the bridging oxygen, this implies that the coverage of  $O_v$  increases due to decrease the bridging oxygen as illustrated in Fig. 8. In noting, when  $(N_{O_b} < N_{O_v})$  deficient the dissociation in  $O_v$  and the defect-free sites, this is in a good agreement with previous works reported in [57]. Exists due to the fact, the dissociation of  $H_2O$  molecules in either terminal hydroxyl or OH in vacancy need a neighbouring oxygen atoms for the H proton creates the second OH hydroxyls on  $O_b$ , besides, the coverage of bridging oxygen decreases with  $H_2O$  increases as Fig. 7(a) improved that clearly. The formation of these radical groups is active to promote the charge separation process as well as the oxidation of organic substances and several applications (self-cleaning, pharmaceutical...) [49] and references therein]. This inhibit the production of OH hydroxyls, meanwhile the molecular adsorption will be the predominant as Fig. 9 illustrates that; where the vacancies increasing the asymmetry for titanium atoms near vacancies that leads to enhance the water adsorption further the  $H_2O$  adsorbed molecular at high and low coverage [38, 53]. From the Eq. (14) it is clearly shown that the coverage of  $Ti_{5c}$  is independent to the flux, it depends only on the concentration of the surface  $O_v$ . Many authors overbalanced the increase to transfer the  $Ti_{6c}$  coordinate underneath the  $O_v$  to  $Ti_{5c}$  [52, 58]. In other hand,  $H_2O$  adsorbed molecular on  $Ti_{5c}$  meanwhile diffuses

toward the vacancy site where the dissociation takes place, thus,  $\text{H}_2\text{O}$  let behind a  $\text{Ti}_{5c}$  site, however, the coverage of Ti increases as illustrated by Fig. 11. Several literature papers conclude that mixed molecular/dissociate is the most stable configuration on rutile  $\text{TiO}_2$  (110) surface [59], besides, the behavior of the water (molecular or dissociative) can change dramatically even for the same materials as the structure and termination of the surface changes.

## 6. Conclusion

The interactions  $\text{H}_2\text{O}$  / rutile  $\text{TiO}_2$  (110) are intriguing subjects; therefore, several experimental and theoretical methods were used to understanding the reaction mechanisms. In terms of theoretical studies, the proposed model presented in this research, a system of coupled differential equations, described and solved in the steady state case taking into account the different reactions of  $\text{H}_2\text{O}$  / rutile  $\text{TiO}_2$  (110). The findings strongly indicate that the interactions  $\text{H}_2\text{O}$ - rutile  $\text{TiO}_2$  (110) have a crucial impact on the production of OH groups at the surface. As one might expect, the coverage of OH hydroxyls increases in the presence of  $\text{O}_V$  which can be applied to enhance the performance of the purification of waste water from organic and inorganic materials plus others applications, but when the concentration of  $\text{O}_V$  increases up to 0.05 leads to reduce the production of OH hydroxyls. This behavior is believed to be a good reason of a modification on surface structure, as well as the  $\text{H}_2\text{O}$  flux. By way of outlook for future challenge given a presence of chemically adsorbed water at rutile  $\text{TiO}_2$  (110) interface and intriguing behavior, this shows the importance of understanding both the structure of rutile  $\text{TiO}_2$  (110) surface and dynamics of  $\text{H}_2\text{O}$  molecules on them. Indeed, knowing the mechanism behavior of water molecules on metal oxides under deferent conditions is very important due to its applications in catalysis.

## Nomenclature

### variables

$\text{H}_2\text{O}$	Water
$\text{Ti}_{5c}$	Titanium five coordinate
$\text{Ti}_{6c}$	Titanium six coordinate
$\text{O}_{2c}$	Oxygen two coordinate (bridging)
$\text{O}_{3c}$	Oxygen three coordinate
$\text{O}_V$	Oxygen Vacancy
OH	Hydroxyl group
$\Phi$	Water flux density per surface site
$\theta$	Coverage
$\theta_{\text{Ti}}$	Coverage of undercoordinated titanium
$\theta_b$	Coverage of oxygen bridging
$\theta_v$	Coverage of oxygen vacancies
$\theta_{\text{H}_2\text{O}}$	Coverage of water molecules on surface
$\theta_{\text{OH}}$	Coverage of hydroxyls groups
$K_i$	The rate constant ( $\text{s}^{-1}$ )
$K_0$	Attempt frequency ( $\text{s}^{-1}$ )
$E_a$	Activation energy (Kj/mole) or (eV)
$T$	Temperature (K)
$R$	Universal gas constant (K j/mole. K)
$ML$	Monolayer
$n_a$	The concentration of the adsorption sites
$n_s$	The concentration of all the atoms in the surface

## Acknowledgement

We gratefully acknowledge financial support of this work from the ministry of higher education and scientific research in Algeria (MESRS).

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