

# Paired Electrochemical Processes: Overview, Systematization, Selection Criteria, Design Strategies, and Projection

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**Abstract:** In most commercial electrochemical processes, only one of the two sides of the cell produces a desirable reaction, whereas the complementary reaction at the other side simply avoids interference with the target process and/or offers fast kinetics as not to limit it. An overview of several possible paired reaction schemes and their advantages is given, accompanied with specific thermodynamic and kinetic selection criteria, cell and reaction design strategies, and examples from the organic, inorganic, and environmental areas. The unnecessary nomenclature dispersion in this field (e.g., combined, concurrent, conjugate, convergent, coupled, divergent, dual, duet, linear, paired, parallel, simultaneous and synchronous electrochemical processes) is revised and a simple, consistent classification, as well as a homogeneous schematization system, is proposed.

**Keywords:** Combined processes; Coupled processes; Electrochemistry; Paired processes; Simultaneous processes; Linear processes; Convergent processes; Divergent processes.

**Resumen:** En la mayoría de los procesos electroquímicos comerciales solo uno de los dos lados de la celda genera un producto de interés, mientras que en el otro lado simplemente se evita que la reacción complementaria interfiera con la reacción principal o bien solo ofrece una cinética que no limite el proceso. Aquí se da una visión global de esquemas posibles de reacciones en pares, así como sus ventajas, acompañada de criterios termodinámicos y cinéticos de selección, estrategias para el diseño de celdas y procesos, y ejemplos de las áreas de orgánica, inorgánica y ambiental. Se analiza la dispersión innecesaria de nomenclatura que existe en la literatura (e.g., procesos electroquímicos combinados, concurrentes, conjugados, convergentes, acoplados, divergentes, duales, en dueto, lineales, en pares, paralelos, simultáneos, y sincronizados) y se propone una clasificación sencilla y consistente así como una homogenización de su esquematización.

**Palabras clave:** Procesos combinados; Procesos acoplados; Electroquímica; Procesos en pares; Procesos simultáneos; Procesos lineales; Procesos convergentes; Procesos divergentes.

## 1. Introduction

Electrochemical industries are among the largest electricity consumers in the manufacturing world. It is still rather common to find processes in which only one of the two sides of the cell produces a desirable transformation, whereas the complementary reaction at the other side is simply selected or allowed as to avoid interference with the target process and/or to offer fast kinetics as not to limit the main activity [1]. This is typically accompanied by a lack of an explicit interest about the use or the commercial value of the products or transformations generated through this complementary process that is usually the decomposition of the solvent. Thus, there is much room for improving the optimal energy expenditure by designing processes where current is made useful at both sides of the cell [2-9]. Even though this in itself does not ensure the total greening of a process, it does contribute towards this objective by responding to one of the main principles of green chemistry: minimize energy requirements [10]. In the present paper, we give an overview of several possible reaction schemes for this to happen and supplement them with specific selection criteria and examples. In addition, the extensive nomenclature

dispersion in this area is revised and a consistent as well as simplified classification and schematization system is proposed.

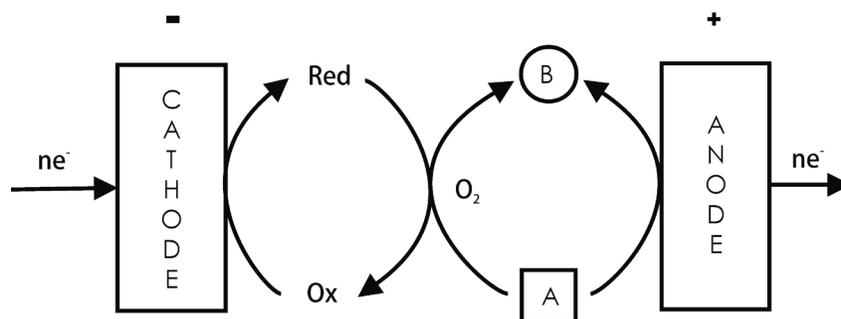
## 2. Paired processes according to their mode of action

We propose that most of the paired processes summarized in the systematic review by Paddon et al. [2] can in principle be grouped according to their *mode of action*. In the following figures a letter inside a square symbolizes a reagent whereas a letter inside a circle symbolizes a product; Red and Ox symbolize the reduced and oxidized species of a redox pair, respectively; and the not enclosed letters symbolize intermediate species. Note that redox mediators reacting *directly* at the electrodes can be designated as *electrochemical mediators*, whereas those acting *in the electrolytic solution* are *homogeneous redox mediators*.

### 2.1 Use of reversible redox mediators

#### 2.1.1 Single redox mediated systems.

An example is the synthesis of *p*-benzoquinone and hydroquinone from benzene [2,11]. The redox mediator in the cathodic

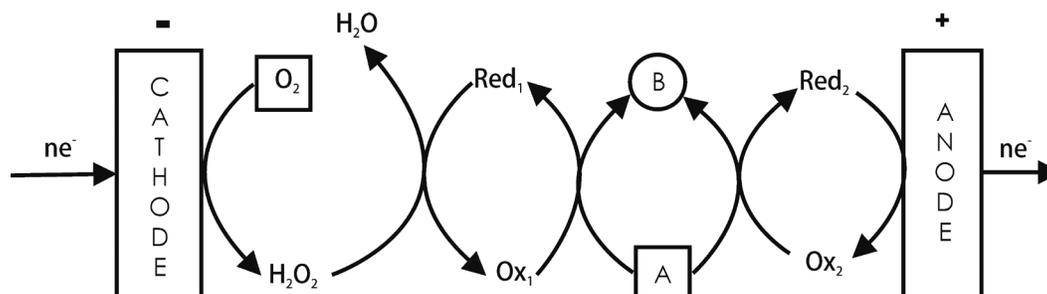


**Fig. 1.** A single redox mediated system. A = benzene, B = hydroquinone + quinone, Red =  $\text{Cu}^+$ , Ox =  $\text{Cu}^{2+}$  [2,11].

side is the  $\text{Cu}^{2+}/\text{Cu}^+$  system.  $\text{Cu}^+$  reduces  $\text{O}_2(\text{g})$  to a peroxy intermediate thus generating a peroxide complex with copper that oxidizes the aromatic ring and yields the same product as in the corresponding anodic oxidation of the aromatic species at a  $\text{PbO}_2$  electrode in a single H-cell (i.e., hydroquinone or quinone). In the paired process, both reactions occur simultaneously in the undivided cell (see Fig. 1). Peroxy intermediates are useful oxidizers produced by cathodic reactions, and whose chemical behavior is equivalent to the anodic reaction as shown in this example, whereby benzene is *electrochemically* oxidized and, at the same time, it is *chemically* oxidized to hydroquinone by this Cu-peroxy intermediate produced cathodically. Depending on experimental conditions, hydroquinone can be oxidized to quinone by the Cu-peroxy complex or else anodically. The result is a paired process with only one final product.

### 2.1.2 Multiple redox mediated systems

An example is the synthesis of sulfones ( $\text{R-SO}_2\text{-R}$ ) from sulfides ( $\text{R-S-R}$ ) [2,5,12]. The redox mediator at the cathodic side is the  $\text{HWO}_5^{2-}/\text{HWO}_4^{2-}$  system, while the  $\text{OsO}_4(\text{OH})_2^{2-}/\text{OsO}_2(\text{OH})_4^{2-}$  system mediates the electron removal at the anodic side. See Fig. 2. As in the previous example, hydrogen peroxide (cathodically produced by the  $2e^-$  dioxygen reduction) is used as the oxidant for generating the tungsten oxidizing species that produce the desired sulfones.



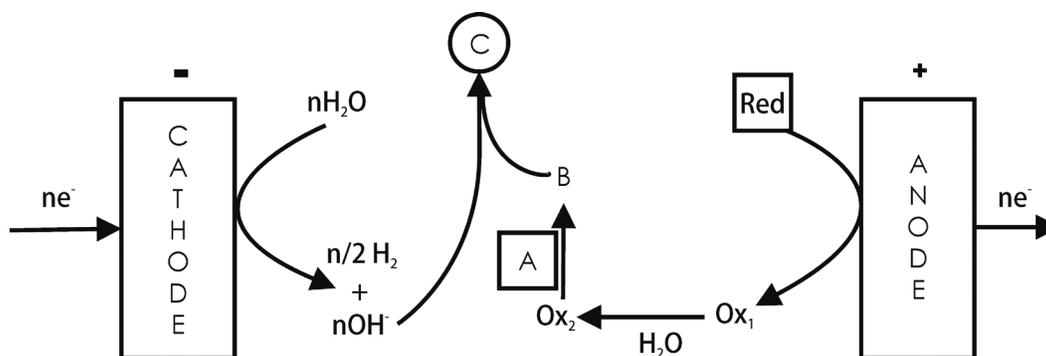
**Fig. 2.** A multiple redox mediated system. A =  $\text{R-S-R}$  (sulfide, where  $\text{R} = \text{CH}_2\text{CH}_2\text{COOH}$  or  $\text{CH}_2\text{CH}_2\text{OH}$ ), B =  $\text{R-SO}_2\text{-R}$  (sulfone), Red<sub>1</sub> =  $\text{HWO}_4^{2-}$ , Ox<sub>1</sub> =  $\text{HWO}_5^{2-}$ , Red<sub>2</sub> =  $\text{OsO}_2(\text{OH})_4^{2-}$ , Ox<sub>2</sub> =  $\text{OsO}_4(\text{OH})_2^{2-}$  [2,5,12].

### 2.2 Use of at least one irreversible redox mediator

An example is the epoxidation of propylene at interdigitated electrodes [2,13]. The redox mediator here is the  $\text{Br}_2/\text{Br}^-$  system. Upon alkaline hydrolysis the anodically made  $\text{Br}_2$  yields  $\text{HBrO}$  that reacts with the double bond of propylene (A =  $\text{CH}_3\text{CH}=\text{CH}_2$ ) to produce the corresponding bromohydrin (B =  $\text{CH}_3\text{CHOHCH}_2\text{Br}$ ). Removal of the bromine atom and the acidic H in the OH- group by the  $\text{OH}^-$  ions coming from the cathodic reduction of water promotes bromine elimination during the 3-member ring closure with the generation of the epoxide. See Fig. 3. The irreversibility stems from side reactions, one of which irreversibly generates  $\text{BrO}_3^-$  ions [13].

### 2.3 Use of at least one reversible and one irreversible redox mediators

An example is the synthesis of epoxides and dibromides from olefins [2,14], where anodically generated  $\text{Br}_2$  adds irreversibly to an olefin (A =  $\text{RCH}=\text{CHR}'$ ) producing the corresponding dibromide (B =  $\text{RCHBrCHR}'$ ). The same olefin is simultaneously oxidized by the redox mediator  $[\text{VO}(\text{acac})_2\text{OOH}]$  to yield its epoxide (C =  $\text{RCHCHR}'\text{O}$ ). The spent homogeneous redox agent is then regenerated by  $\text{H}_2\text{O}_2$  produced cathodically from  $\text{O}_2$ . See Fig. 4.



**Fig. 3.** An irreversible redox mediated system. A =  $\text{CH}_3\text{CH}=\text{CH}_2$ , B =  $\text{CH}_3\text{CHOHCH}_2\text{Br}$ , C =  $\text{CH}_3\text{CHCH}_2\text{O}$ , Red =  $\text{Br}^-$ ,  $\text{Ox}_1 = \text{Br}_2$ ,  $\text{Ox}_2 = \text{BrOH}$  [2,13].

#### 2.4 Use of multiple electron exchange processes

An example is the cathodic generation of nitrogen heterocyclic anions that react with anodically generated radical aromatic cations [2,15]. For example, 1 *H*-tetrazole (the simplest tetrazole, see species A =  $\text{CH}_2\text{N}_4$  in Fig. 5) is cathodically converted by a one-electron transfer into its anion B ( $\text{CHN}_4^-$ ), which then pairs in a nucleophilic fashion with the electron-rich cation D ( $p\text{-(CH}_3\text{O)}_2\text{C}_6\text{H}_4^+$ ) produced by the anodic oxidation of 1,4-dimethoxybenzene (C =  $p\text{-(CH}_3\text{O)}_2\text{C}_6\text{H}_4$ ). This step generates a complex intermediate that releases a second electron at the anode, yielding a mixture of tetrazolemethoxyarenes. See Fig. 5.

#### 2.5 Use of multiple chemical reaction processes coupled to electron exchange processes

An example is the formation of protected homoallylic alcohols (see Fig. 6). Here, the cation F ( $\text{CH}_2=\text{N}^+(\text{CH}_3)\text{CHO}$ , generated from the oxidation of *N,N*-dimethylformamide used as solvent) reacts with the anion D ( $\text{CH}_2=\text{CHCH}_2\text{CH}(\text{Ar})\text{O}^-$ ) resulting from the reaction between an aromatic aldehyde (C =  $\text{ArCHO}$ ) and the allylic anion B ( $\text{CH}_2=\text{CHCH}_2^-$ ), generated by the cathodic dehalogenation of allyl bromide, A ( $\text{CH}_2=\text{CHCH}_2\text{Br}$ ) [2,16]. Both electron-transfer processes at the electrodes are

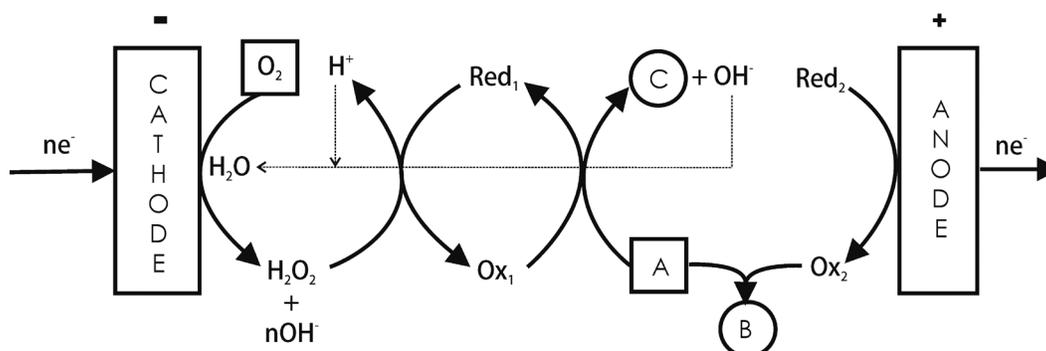
followed by consecutive chemical reactions that yield the final product, G ( $\text{CH}_2=\text{CHCH}_2\text{CH}(\text{Ar})\text{OCH}_2\text{N}(\text{CH}_3)\text{CHO}$ ). The cationic intermediate of DMF F ( $\text{CH}_2=\text{N}^+(\text{CH}_3)\text{CHO}$ ) is obtained by a two-electron, one-proton release from the solvent, whereas the alkoxy intermediate D ( $\text{CH}_2=\text{CHCH}_2\text{CH}(\text{Ar})\text{O}^-$ ) is obtained after a 2-electron incorporation accompanied by bromide elimination to obtain the stabilized allylic anion.

### 3. Terms used in the literature

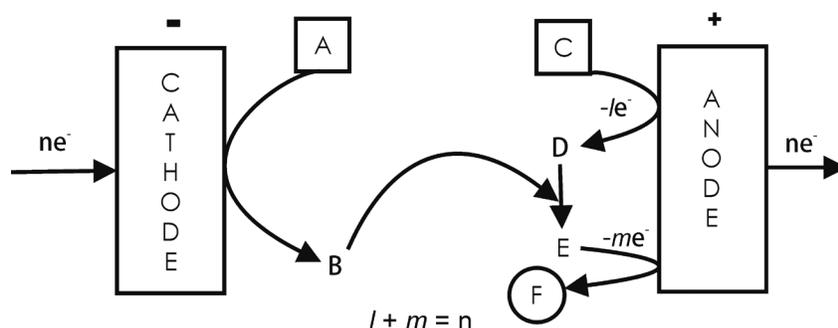
All of these processes involve steps that are more elementary and/or can be classified in global, simpler schemes. We have identified all the terms devoted in the literature to designate that useful reactions occur at both sides in an electrochemical cell; they are rather diverse and do not necessarily convey similar meanings, which leads to confusion and scattering. Table 1 summarizes the various terms that have been used.

### 4. Classification proposal

As can be easily perceived, this great disparity in terms and meanings obstructs communication and understanding within



**Fig. 4.** A system using at least one reversible and one irreversible redox mediators. A =  $\text{RCH}=\text{CHR}'$ , B =  $\text{RCHBrCHBrR}'$ , C =  $\text{RCHCHR}'\text{O}$ , Red<sub>1</sub> =  $[\text{VO}(\text{acac})_2]$ , Ox<sub>1</sub> =  $[\text{VO}(\text{acac})_2\text{OOH}]$ , Red<sub>2</sub> =  $\text{Br}^-$ , Ox<sub>2</sub> =  $\text{Br}_2$  [2,14].



**Fig. 5.** A system using two or more electron exchange processes. A =  $\text{CH}_2\text{N}_4$ , B =  $\text{CHN}_4^-$ , C =  $p\text{-(CH}_3\text{O)}_2\text{C}_6\text{H}_4$ , D =  $p\text{-(CH}_3\text{O)}_2\text{C}_6\text{H}_4^+$ , E = 1,4-dimethoxybenzene-tetrazole complex ( $\text{C}_8\text{H}_{10}\text{O}_2\text{CHN}_4$ ), F = a mixture of isomeric tetrazole-1,4-dimethoxybenzenes,  $\text{C}_8\text{H}_9\text{O}_2\text{CHN}_4$ ,  $l = 1$ ,  $m = 1$ ,  $n = 2$  [2,15].

the electrochemical community and with people in other fields as well. Consequently, a simplified global scheme that encompasses all the terms included in Table 1 should be based on: a) reserving the terms *combined* and *coupled* to their most common usage on this table, b) noting that the terms *concurrent*, *conjugate*, *dual*, *paired*, and *simultaneous* refer essentially to the same type of processes, and c) noting that the term *duet* has the same meaning as the term *linear* and corresponds to processes where the same product is generated at both sides of an electrochemical cell through the production of reactive redox mediators at the electrodes. Therefore, these reactions can be clustered under the term *linear paired electrolysis* [51].

All these names of reactions can be accommodated within the term used by most references (i.e., *paired*). Therefore, we hereby match this proposal with the *simplified* global nomenclature system and schemes for the paired electrochemical reactions essentially proposed by Aust [3], Steckhan [42], and Paddon et al. [2] and merged as follows.

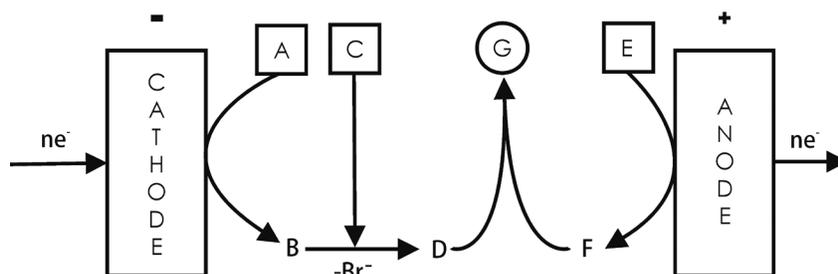
#### 4.1 Parallel paired electrolysis

Both cathodic and anodic reactions are matched and contribute usefully to the formation of the final product(s) (see Fig. 7) [2]. Examples include the industrial cathodic synthesis

of phthalide (B =  $\text{C}_6\text{H}_4\text{COOCH}_2$ ) from dimethyl phthalate (A =  $\text{C}_6\text{H}_4(\text{C}_2\text{H}_3\text{O}_2)_2$ ) with the simultaneous anodic production of *tert*-butyltoluene dimethylacetal (D =  $(\text{CH}_3)_3\text{CC}_6\text{H}_4\text{CH}(\text{OCH}_3)_2$ ) from *tert*-butyltoluene (C =  $(\text{CH}_3)_3\text{CC}_6\text{H}_4\text{CH}_3$ ) [3,7]; the cathodic synthesis of aniline from nitrobenzene with the simultaneous anodic generation of  $\text{Cl}_2$  from NaCl [25]; and the galvanostatic synthesis of Michael adducts (an enolate in the cathodic side and the Lewis acid  $\text{BF}_3$  in the anodic side, in the ionic liquid 1-ethyl-3-methylimidazolium tetrafluoroborate) [52].

#### 4.2 Convergent paired electrolysis

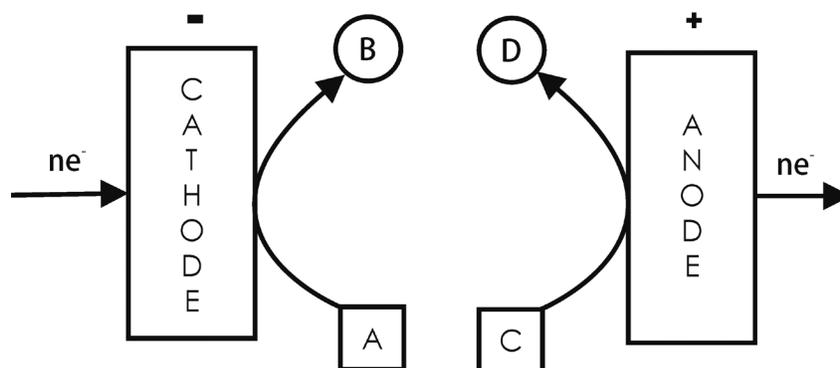
The products generated at both sides of the cell react to form one final product (see Fig. 8) [3]. Examples include the generation of  $\text{ClO}_2$  from  $\text{ClO}_3^-$  (at the cathode) and from  $\text{ClO}_2^-$  (at the anode) [53]; the production of simultaneous electroluminescence from the emission of the 3-aminophthalate dianion radical (generated at the cathode by oxidation performed by the hypochlorite ions generated from the reduction of  $\text{ClO}_2^-$ , and at the anode by the direct oxidation of luminol) [54]; and the simultaneous production of glyoxalic acid (C =  $\text{CHOCOOH}$ ) by the cathodic reduction of oxalic acid (A =  $\text{COOHCOOH}$ ) and the anodic oxidation of glyoxal (B =  $\text{CHOCHO}$ ) [55,56].



**Fig. 6.** A system using two or more chemical reaction processes coupled to the electrochemical processes. A =  $\text{CH}_2=\text{CHCH}_2\text{Br}$ , B =  $\text{CH}_2=\text{CHCH}_2^-$ , C = ArCHO, D =  $\text{CH}_2=\text{CHCH}_2\text{CH}(\text{Ar})\text{O}^-$ , E =  $(\text{CH}_3)_2\text{NCHO}$ , F =  $\text{CH}_2=\text{N}^+(\text{CH}_3)\text{CHO}$ , G =  $\text{CH}_2=\text{CHCH}_2\text{CH}(\text{Ar})\text{OCH}_2\text{N}(\text{CH}_3)\text{CHO}$ . Here, Ar =  $\text{C}_6\text{H}_5$ , 4-MeC<sub>6</sub>H<sub>4</sub>, or 4-MeOC<sub>6</sub>H<sub>4</sub> [2,16].

**Table 1.** Terms employed in the electrochemical literature to denote different types of paired reactions.

TERM	LITERATURE USAGE	REFERENCE
Combined electrochemical process	Electrochemical methods are combined with other electrochemical or non-electrochemical treatments to eliminate persistent pollutants.	[16,17,18,19]
Idem	A desired substance is produced with the cogeneration of a useful byproduct.	[20]
Idem	Useful processes occur at both electrodes.	[17]
Idem	In a galvanic process, the products of both electrodic reactions combine to yield a substance that generates useful chemistry.	[21]
Concurrent electrochemical process	Useful processes occur at both electrodes.	[22,23]
Conjugate electrochemical synthesis	Useful products are generated at both electrodes.	[24]
Convergent electrosynthesis	The same product is generated from the same starting material both in the anodic and cathodic chambers.	[2]
Convergent paired electrosynthesis	The products generated from both electrodes react to form one final product.	[25]
Idem	Both cathodic and anodic reactions contribute usefully to the formation of a final product(s).	[26]
Idem	Two different starting materials yield the same product at each electrode.	[2]
Idem	The same product is synthesized by the anodic and cathodic processes.	[3]
Coupled electrochemical reactions	Electrochemical or chemical processes are closely associated with the main electron transfer in an electrochemical reaction, e.g., EC, CE, ECE, etc.	[27-29]
Idem	A substance produced in one electrodic chamber is transported and reacted at the other chamber or electrode.	[30-34]
Divergent paired electrosynthesis	Two different products are formed at the electrodes using a single starting material.	[2]
Idem	Using one single substrate creates an oxidized and a reduced product.	[3]
Dual electrolysis	Useful intermediates are generated at both electrodes.	[35]
Duet electrosynthesis	The same product is generated from a single starting material in both the anodic and cathodic chambers in a single electrolytic cell.	[36,37]
Linear paired electrosynthesis	One product is synthesized from the substrate by one or two-mediated electrode processes.	[3]
Paired electrosynthesis	Both cathodic and anodic reactions are matched and contribute usefully to the formation of the final product(s).	[2]
Paired electrolysis	Both cathodic and anodic reactions contribute usefully to the formation of the final product(s).	[9,38]
Idem	A substance produced in one electrodic chamber is transported and reacted at the other chamber or electrode.	[38,39]
Idem	Intermediate products formed at the anode and cathode react with the same substance to form the final product(s).	[40]
Idem	Two starting materials are converted into two useful products without any chemical interaction during the electrolysis process.	[41]
Idem	Both cathodic and anodic reactions generate high-value product(s).	[7,42]
Parallel paired electrosynthesis	Idem	[2]
Parallel electrosynthesis	The products generated at the anode and cathode do not interfere nor react.	[43,44]
Simultaneous electrochemical processes	A substance produced in one electrodic chamber is transported and reacts at the other chamber or electrode.	[45]
Idem	Useful processes occur at both electrodes.	[46-49]
Idem	A photogalvanic process generates electricity and a useful product at one electrode, and a useful electrochemical transformation at the other.	[50]



**Fig. 7.** Parallel paired electrolysis [3,7,24]. Cathodic synthesis of phthalide from dimethyl phthalate with the simultaneous anodic production of *tert*-butyltoluene dimethylacetal from *tert*-butyltoluene: A =  $C_6H_4(C_2H_3O_2)_2$ , B =  $C_6H_4COOCH_2$ , C =  $(CH_3)_3CC_6H_4CH_3$ , D =  $(CH_3)_3CC_6H_4CH(OCH_3)_2$  [7].

#### 4.3 Divergent paired electrolysis

Two different products are formed at the electrodes using a single starting material (see Fig. 9) [2]. Examples include the production of sorbitol (B =  $CH_2OH(CHOH)_4CH_2OH$ ) at the cathode and of gluconic acid (C =  $CH_2OH(CHOH)_4COOH$ ) at the anode from glucose (A =  $CH_2OH(CHOH)_4CHO$ ) [57-59]; the generation of *l*-cysteic acid (at the anode) and *l*-cysteine (at the cathode) from *l*-cystine [60,61]; the cathodic generation of ethylene and the anodic generation of oxalic acid from acetylene [62]; and the simultaneous anodic acetoxylation and cathodic carboxylation of conjugated dienes (e.g., 1,3-cyclohexadiene) and  $CO_2$  to produce the corresponding 1,4-diacids and 1,4-diols used as key synthetic precursors [63].

#### 4.4 Linear paired electrolysis

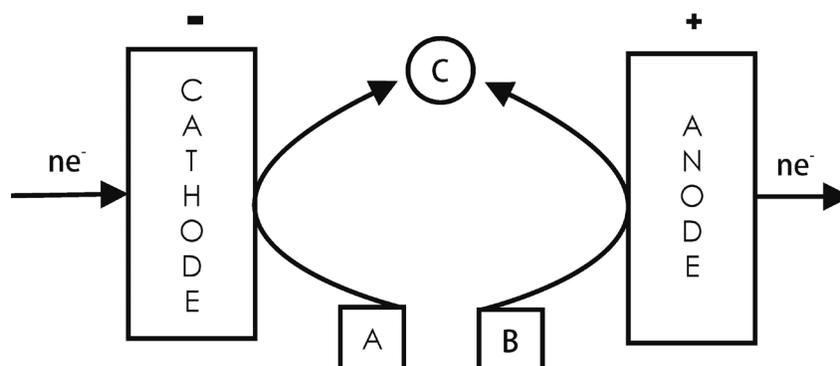
One product is synthesized at both sides from the substrate by one or two-mediated electrode processes (not shown in this figure, see the *simplified* Fig. 10) [3]. For a more complete scheme, see for example Fig. 2. Examples include the production of *n*-butyric acid from *n*-butanol through its indirect anodic oxidation (by  $NiO(OH)$ ) and indirect cathodic oxidation (by  $H_2O_2$ ) [64]; the simultaneous oxidation of sodium gluconate to D-arabinose

using the  $Fe^{3+}/Fe^{2+}$  and  $ClO^-/Cl^-$  redox mediators in the anolyte and the  $Fe^{3+}/Fe^{2+}$  and  $V^{5+}/V^{4+}$  mediators in the catholyte [65]; the indirect cathodic oxidation of a hydroxylamine (A =  $CH_3(CH_2)_3N(OH)(CH_2)_3CH_3$ ) by the  $O_2/H_2O_2/WO_5^{2-}/WO_4^{2-}$  redox system to produce the corresponding *N*-amine *N*-oxide (i.e., a nitrene, B =  $CH_3(CH_2)_3N(O)=CH(CH_2)_2CH_3$ ), coupled to the indirect anodic oxidation of the same hydroxylamine in the presence of a halide/halogen system (i.e.,  $Br_2/Br^-$  or  $I_2/I^-$ ) [12]; and the production of amino imino methane-sulfonic acids from the corresponding thioureas, catalyzed by the  $WO_5^{2-}/WO_4^{2-}$  and  $Cl_2/Cl^-$  redox pairs. Here, the  $WO_5^{2-}$  is continuously regenerated by  $H_2O_2$  coming from the cathodic reduction of  $O_2$ , as shown in Fig. 2. The nature of the sulfur-containing group defines the type of current efficiency response with the applied current density [66].

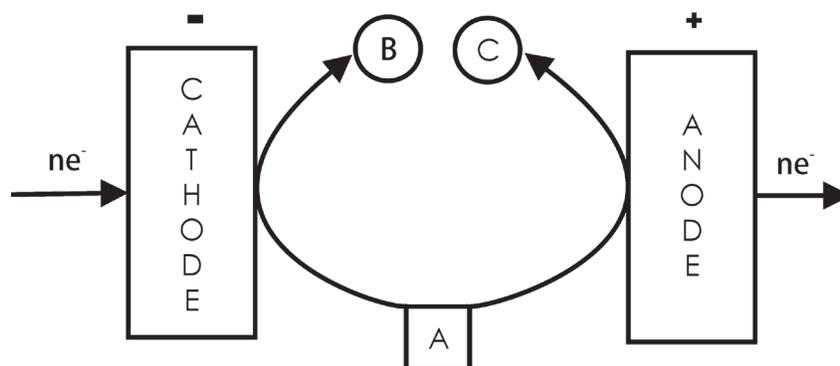
The paired processes discussed so far, essentially involve organic starting materials and products (other than the redox mediators). Particular processes in other areas are presented next.

### 5. Inorganic paired electrochemical processes

Compared to the achievements and proposals in the organic paired electrosynthesis field, inorganic processes are rather



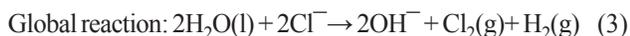
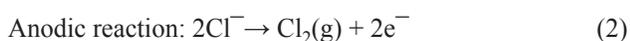
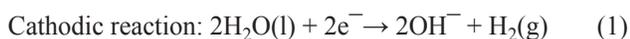
**Fig. 8.** Convergent paired electrolysis [53-56]. Simultaneous production of glyoxalic acid by the cathodic reduction of oxalic acid and the anodic oxidation of glyoxal: A =  $COOHCOOH$ , B =  $CHOCHO$ , C =  $CHOCOHOH$  [55,56].



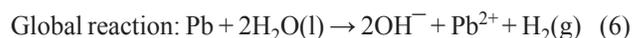
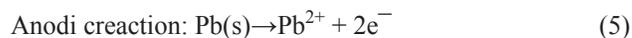
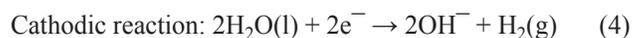
**Fig. 9.** Divergent paired electrolysis [57-62]. Production of sorbitol (at the cathode) and of gluconic acid (at the anode) from glucose: A =  $\text{CH}_2\text{OH}(\text{CHOH})_4\text{CHO}$ , B =  $\text{CH}_2\text{OH}(\text{CHOH})_4\text{CH}_2\text{OH}$ , C =  $\text{CH}_2\text{OH}(\text{CHOH})_4\text{COOH}$  [57-59].

scarce [67]. Selected examples are summarized below. (For simplicity, the physical state of aqueous species is omitted.)

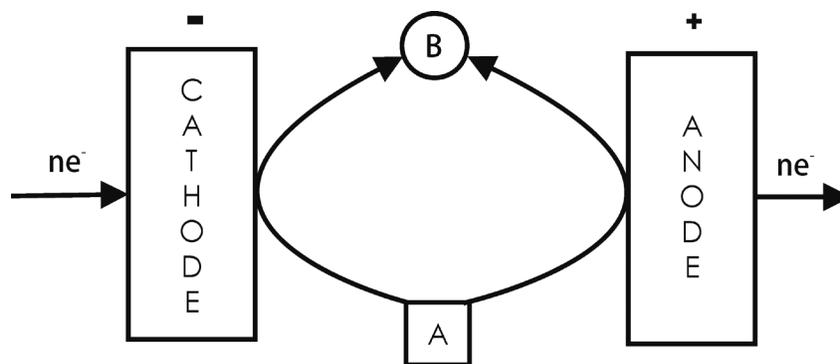
**5.1** The well-known chlor/alkali process is the most important paired process from the commercial standpoint. World production of these two inorganic compounds (i.e.,  $\text{Cl}_2$  and  $\text{NaOH}$ ) is the highest among all the electrochemical processes and it is higher than almost all the inorganic chemical processes (except that of sulfuric acid). The cathodic reaction consists of water reduction, whereas the anodic reaction consists of the oxidation of chloride ions from brine as the raw material [68]. This is a parallel paired electrolysis (see Fig. 7). The reactions involved in the process are:



**5.2** In a cell equipped with a Pb anode and a  $\text{Na}_2\text{CrO}_4$  aqueous solution in the anolyte,  $\text{PbCrO}_4(\text{s})$  is produced in this chamber while  $\text{NaOH}$  is generated in the catholyte as a result of  $\text{Na}^+$  migration and  $\text{OH}^-$  production from water electrolysis. Linear voltammetry signaled the appropriate potential to be used for Pb dissolution. Cell voltage vs. time curves helped to identify undesired electrode passivation by an insoluble product layer [69]. This is an example of parallel paired electrolysis (see Fig. 7).

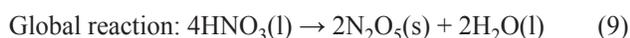
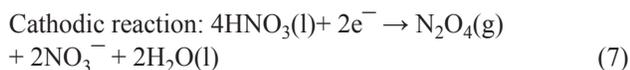


**5.3** An unusual oxidant production scheme involves dehydration of  $\text{HNO}_3$  to  $\text{N}_2\text{O}_5$  (note the same oxidation state for N in both compounds), where  $\text{HNO}_3$  is reduced at the cathode to yield  $\text{N}_2\text{O}_4$ . Due to the high volatility of the latter, it is then

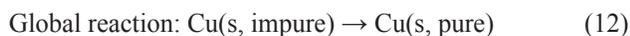
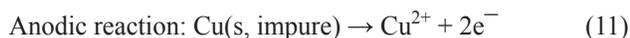
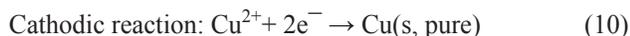


**Fig. 10.** Simplified linear paired electrolysis scheme [2,3,12,62,63,65,66]. A more complete description of these systems can be seen in Fig. 2. For example, a) A = *n*-butanol, B = *n*-butyric acid [64]; b) A = sodium gluconate, B = D-arabinose [65]; c) A =  $\text{CH}_3(\text{CH}_2)_3\text{N}(\text{OH})(\text{CH}_2)_3\text{CH}_3$ , B =  $\text{CH}_3(\text{CH}_2)_3\text{N}(\text{O})=\text{CH}(\text{CH}_2)_2\text{CH}_3$  [12]; and d) A = a thiourea, B = the corresponding amino imino methane-sulfonic acid [63].

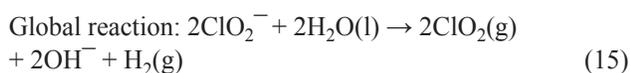
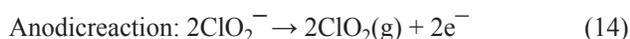
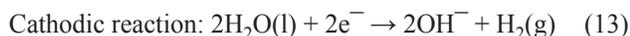
distilled from the cooled catholyte and directed towards the anolyte where re-oxidation occurs to generate  $\text{N}_2\text{O}_5$  (i.e., a much safer nitrating agent than concentrated  $\text{HNO}_3$ ) [68,70]. An alternative is to avoid the heating step by using a membrane process [71]. This is a parallel paired electrolysis (see Fig. 7); in this example, product B from the cathode becomes reagent C in the anolyte.



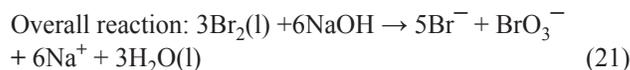
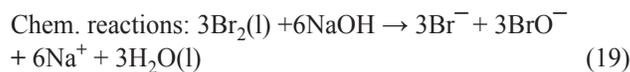
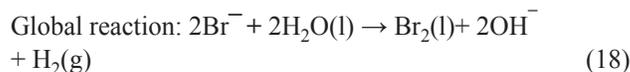
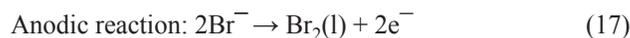
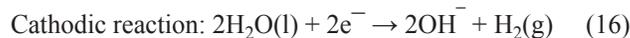
**5.4** In the classical process of electrorefining of metals applied on a large scale to the electrolytic refinement of copper, impure copper anodes are dissolved electrolytically to produce copper (II) ions, which are then transported to the cathode whereby they deposit without the impurities contained in the initial copper metal source [72]. This is a parallel paired electrolysis (see Fig. 7), except that in this example product D from the anode becomes reagent A in the catholyte.



**5.5** The co-production of  $\text{ClO}_2$  and  $\text{NaOH}$  is achieved in a divided cell by the anodic oxidation of a  $\text{NaClO}_2$  solution with simultaneous  $\text{H}_2\text{O}$  reduction. The  $\text{Na}^+$  ions migrate towards the cathode through a cation exchange membrane to compensate for the negative charges of the  $\text{OH}^-$  generated there and produce  $\text{NaOH}$  [73]. This is a parallel paired electrolysis (see Fig. 7).



**5.6** The production of the oxidizing agent  $\text{NaBrO}_3$  is based on the initial anodic oxidation of  $\text{Br}^-$ . The  $\text{Br}_2$  thus produced reacts with  $\text{NaOH}$  generated at the cathodic side from water electrolysis, giving rise to a disproportionation reaction that produces a mixture of  $\text{Br}^-$  and  $\text{BrO}^-$  ions. Afterwards, three  $\text{BrO}^-$  ions disproportionate yielding the desired product, i.e.,  $\text{BrO}_3^-$  and two  $\text{Br}^-$  ions that can in principle be recycled into the process [74]. This is a parallel paired electrolysis (see Fig. 7) except that in this example products B and D react to form  $\text{BrO}^-$  that disproportionates as described above.



## 6. Environmentally-related paired electrochemical processes

Fortunately, the idea of paired processes has reached the environmental arena and some processes have been developed in this area as well. Selected examples are described next. Since the objective of most of them is not a synthetic one, the processes in this section are different from those described above throughout this paper.

### 6.1 Paired organic pollutant degradation

A proof-of-concept of the paired electrochemical treatment of two organic molecules involves the cathodic reduction of dichlorophenol and the anodic oxidation of phenol in aqueous solution [75]. Another example involves the cathodic production of hydrogen peroxide from dioxygen reduction with the simultaneous degradation of phenol, where both processes occur on boron-doped diamond electrodes. These processes are carried out at the same acidic pH in order to use the peroxide produced in a second pollutant degradation scheme using Fenton's reaction (see below) [76].

### 6.2 Paired ligand degradation and metal ion recovery from metal complexes

The cathodic removal of metal ions (e.g.,  $\text{Cd}(\text{II})$ ,  $\text{Cu}(\text{II})$ ,  $\text{Ag}(\text{I})$ ) and the anodic oxidation of their ligands (typically  $\text{CN}^-$ ) from the electroplating, metal finishing, and metallurgical industries have been thoroughly studied [30-32, 46, 77]. While the chemical route generates large amounts of sludge containing metal hydroxides [31], the electrolytic route offers the possibility of recovering metals in their elementary solid form (or else as oxides [47]), while directly or indirectly destroying their ligands [30,31,46,48,77]. In the case of  $\text{CN}^-$  ligands, such destruction is through their oxidation to  $\text{CNO}^-$  (which is much less toxic than  $\text{CN}^-$ ) followed by further oxidation to innocuous  $\text{N}_2$ . Seawater may be added to the process to generate chlorinated

oxidants capable of performing this transformation into  $N_2$  [46,48]. The undesirable side oxidation of  $CNO^-$  into  $NO_3^-$  ions can be minimized by regulating the electrolysis time [78]. Interestingly, a judicious combination of surface area (that controls the potential in a typical galvanostatic process) and pH can facilitate the global reduction of  $NO_3^-$  to  $N_2(g)$  [45].

### 6.3 Paired COD (chemical oxygen demand) removal and metal ion recovery

In a similar fashion as above, the paired cathodic removal of copper ions and the anodic oxidation of COD can be achieved with excellent yields [30,33,47,79]. Part of this success has been attributed to the simultaneous cathodic generation of strong oxidizers capable of attacking organic matter [33].

### 6.4 Paired inorganic pollutant degradation or removal

A pollutant is transformed into its pure, useful elemental components in very few cases. In one of them, the highly noxious  $H_2S$  is electrochemically treated either directly or indirectly with the resulting simultaneous production of high purity  $H_2(g)$  at the cathode and of  $S(s)$  in the anolyte ( $H_2SO_4$  or  $IO_3^-$  ions can also be produced) [80-82]. In the indirect case, the gas is first absorbed into a liquid containing an oxidizing agent that oxidizes  $S^{2-}$  ions to elemental  $S(s)$ . The spent oxidizer can be anodically regenerated and returned to the absorbing system in an outer-cell process [34,83]. Simultaneously, protons from water are reduced at the cathode and produce  $H_2(g)$ . Electrochemical methods offer an additional variety of possibilities for  $H_2S$  treatment, although not all of them are paired processes [83-92]. Another process involves the electrochemical treatment of N-containing wastes by the paired cathodic reduction of  $NO_3^-$  ions and the anodic oxidation of  $NH_3$  in aqueous solution. This can be achieved using a zero-gap solid polymer electrolyte, *ZGSPE*. Although the selectivity for the  $NH_3$  oxidation to  $N_2$  needs enhancement, no undesirable  $NO_2^-$  ions are generated and  $N_2$  is the main product [39]. In yet a different case, Ag can be cathodically recovered from photographic wastes containing  $Ag^+$  and  $S_2O_3^{2-}$  ions. Here, the anodic generation of  $H^+$  provides the required low pH for the disproportionation of  $S_2O_3^{2-}$  into  $S(s)$  and  $SO_3^{2-}$  or  $SO_2(g)$  [51], or else into  $S^{2-}$  and  $SO_4^{2-}$  [93]. The sulfide ions then precipitate with the silver (I) ions present, forming insoluble  $Ag_2S$  that is recovered easily by filtration. Interestingly, insoluble  $Ag_2S$  is also generated in the catholyte by the reaction between the  $Ag(I)$  ions and the  $S^{2-}$  ions produced by the reduction of  $S_2O_3^{2-}$ , making this a *linear paired process* [51].

### 6.5 Paired organic or inorganic pollutant removal and useful water electrolysis

Flue gas desulfurization (FGD) processes for the removal of  $SO_2(g)$ , are typically based on its *reversible* absorption in sodium citrate solutions [94], or its essentially *irreversible* absorption in basic solutions [95,96]. Problems involving slow kinetics

and abundant sludge generation have promoted the search for alternative remediation methods. Electrochemical schemes have been developed based on the direct or indirect anodic oxidation of  $SO_2(g)$  to  $SO_4^{2-}$  (which in acidic media forms  $H_2SO_4$ ) or its reduction into elemental  $S(s)$ ; the paired reduction of water produces useful  $H_2(g)$  [95,97-101]. In another example, organic textile dye effluents are anodically degraded with the paired cathodic generation of  $NaOH$  [49] or  $H_2(g)$  production [102] in an electrochemical membrane cell. Additionally, a paired electroprecipitation and electrogenerated Fenton oxidation method (called *peroxicoagulation*) for the removal of organics (e.g., vegetal tannins) involves the introduction of  $Fe(II)$  ions from an Fe anode into the solution with the paired cathodic generation of  $H_2O_2$  from  $O_2(g)$  reduction [103,104]. The combination of  $Fe(II)$  with the peroxide generates the well-known Fenton's reagent that produces highly oxidizing  $\bullet OH$  radicals,  $Fe(III)$ , and  $OH^-$  [105]. The mineralization accomplished by the radicals, together with the precipitation of the resulting  $Fe(III)$  oxyhydroxides, is very effective for the removal of pollutants from synthetic wastewaters [17, 106-108].

### 6.6 Paired production of environmentally friendly oxidizing/disinfecting agents

The paired production of  $O_3(g)$  and  $H_2O_2$  in a flow reactor equipped with a proton exchange membrane avoids the use of two separate cells. In the first one,  $O_3(g)$  is anodically produced and  $H_2(g)$  is sacrificially generated at the cathode, whereas in the second one,  $H_2O_2$  is produced at the cathode and  $O_2(g)$  is sacrificially generated at the anode [109]. Another example involves the anodic and cathodic electrogeneration of  $H_2O_2$  and  $ClO^-$  ions from dissolved  $O_2(g)$  and  $Cl^-$  ions, respectively, in aqueous solution. The combined action of both oxidizing agents degrades selected organics (e.g., formaldehyde) very effectively [40]. A third example involves the production of a disinfecting agent at both sides of the cell through the generation of  $ClO_2$  from  $ClO_3^-$  (at the cathode) and  $ClO_2^-$  (at the anode) as discussed above [53].

## 7. Advantages of paired processes

As discussed earlier, multiple organic synthetic routes have been designed using this concept (see for example the reviews in refs. 2-8), which has been extended to other fields, such as inorganic synthesis [8, 67-74] and environmental remediation [35,42,75]. Benefits include not only economic aspects, but also electricity savings that contribute to a decrease in fossil fuel usage with its concomitant combustion waste production. This approach also aims at useful products on both sides circumventing further treatments and costly disposal [3,8]. Other advantages of these processes compared to conventional electrolysis include:

**7.1** Energy consumption can be reduced as much as 50% and time can be saved [35,40,65].

**7.2** Current efficiency can theoretically be increased up to 200% [3,40,41,65]. This is so because current efficiency can be defined as the moles of products per unit of charge moved in the cell [110]. Therefore, yields can be higher than 100% in cells where both sides of the current are used to advantage [2,7,12,14,27,41,66,110,111].

**7.3** Resources are preserved by avoiding the need for additional separate electrodes, electrolytes, etc. [3].

**7.4** Wastes/emissions can be minimized and processes made greener with higher atom economies [3,7,110].

**7.5** Instead of the requirement for two separate reactors, space can be saved and optimized.

**7.6** The economics of a process benefit as fewer separation and synthetic steps are required.

## 8. Reactor/cell design strategies

Due to the highly specific nature of the different processes and the fact that literature reports on paired developments seldom delve into the intricacies of reactor/cell design, a generalization is not possible. However, some observations may offer an initial guide.

**8.1** The issues of product separation and isolation are of fundamental relevance since an intermediate or product must not return to the opposite electrode where it was formed before reacting (i.e., electrochemical short circuits) as this could destroy key components or reverse the desired reaction. This effect, together with possible reactant cross-contamination, would represent a serious economic drawback [8] and must be avoided [2,3,9,101]. Therefore some researchers have used H-type cells [2,11,12,14,37,64,67,111] and other cells with diverse separation strategies (e.g., fritted glass [40,41,52,65,112,113], porous polymeric membranes [58,71], ion exchange membranes [49,101,102,109,114-116], and ion exchange resins [51]). Fitting electrodes directly into membranes permits more compact designs [101]. Membranes also avoid electrolyte exchange generated by pressure variations [101]. Undivided batch [30,47,78,111,117-119] or semi-batch [38], as well as flow-by [120], flow-through [2,58,67,113,121], flow-through coupled to a continuous stirred tank reactor (CSTR) [113], radial flow [48], and thin layer flow cell systems [122,123] have been employed. Some even consist of 3-compartments that prevent electrochemical short circuits by selectively avoiding reactant migration [101,116,124].

**8.2** To determine the current density in an electrolytic reactor design it is necessary to define and/or calculate several engineering parameters like fluid dynamics, heat and mass transfer, separation processes, process modeling and optimization, and process control. As in kinetic models, the determination of

current density is a function of the diffusional and activation current densities. At the same time, the activation parameters are expressed in terms of Tafel behavior [125]. For this reason, the capacitive currents represent a less important design factor concerning the operating current density for different paired reactions. The operational current must be smaller than the limiting current of either side of the cell so that the global process can be made efficient and produce high yields. This is achieved when the mass transfer coefficients are raised by either increasing the parameters associated with the corresponding diffusion coefficient or decreasing the convection layer thickness by increasing the convective term of the mass transfer expression [123]. Current densities may vary widely according to the system (e.g., from 1 to 1000 mA cm<sup>-2</sup> [37,101]).

To obtain the optimum current density it is necessary to maximize its function through an optimization technique as to allow finding the optimum current density without having to resort to heuristic rules, as it was done in the past (generally based on a percentage of the limiting current density) [126]. Simulations and modeling of electrochemical reactors using, for example, the finite element method [127], dynamic simulation [128], and the modified Shooting-Runge Kutta numerical method [113,129] can save time and cost in the design/selection of systems from an engineering point of view.

**8.3** Another key aspect to be evaluated during electrochemical reactor design (and consequently to determine the operating current density) is the cell potential. This is of great practical importance since electricity costs are directly proportional to it. As stated earlier, due to the paired processes characteristics many of them use membranes or diaphragms. In these cases, the parameters to determine the cell potential include the anodic and cathodic potentials, as well as the voltage drops caused by ohmic resistance in the anolyte, catholyte, and in the diaphragm or membrane. To minimize the cell voltage for a given current density it is necessary to: 1) properly select the electrode materials for each reaction; 2) minimize the distance between electrodes; 3) maximize the electrolyte conductivities; and 4) select a membrane (or a diaphragm) that produces the lowest potential drop [125]. Ohmic drop considerations are obviously important especially when using membranes, since for example Nafion 125 can offer a resistance of up to 5 ohms cm<sup>2</sup> at room temperature even in a 55% H<sub>2</sub>SO<sub>4</sub> electrolyte [101].

**8.4** In addition to the ohmic drops it is necessary to determine other reactor parameters such as space-time yield and chemical yield. The former is defined as the mass of product generated by a unit reactor volume per unit of time. The chemical yield is the maximum amount of product that can be obtained for a 100% conversion of the reactant [125].

**8.5** Mass transfer can be enhanced for example by vigorous stirring [30,37,47,67,121,130], gas bubbling [12,14,30,40,64,65,114], or by using tridimensional electrodes (e.g., metal meshes [56], porous [17,57,101,121,131,132],

particulated [57], or fluidized beds [69]). When volatile solvents are used, reflux condensation may be necessary [130].

**8.6** In the case of homogeneous reactions between two short-lived intermediates, microflow systems [4,133,134] and interdigitated band electrode arrays (IBEA) with narrow interelectrode gaps (i.e., 0.25 - 1 mm) that allow an overlap of the cathodic and anodic diffusion layers have been used [2,127,135]. These offer low ohmic losses, enhanced mass transfer, and much higher efficiencies than, for example, parallel plate electrode configurations [127]. Some even avoid the requirement for intentionally added electrolytes [134], thereby increasing the potential for low-emission processes.

## 9. Additional selection criteria for paired reactions related to energy savings

In addition to the considerations discussed above, other factors and recommendations include the following:

**9.1** Increased efficiency should be attempted by using side-by-side reactions, neither of which individually is lower in voltage than commercial half-cell reactions, but which together result in an increased voltage drop and high-energy savings. A wrong approach would be to simply prioritize lower voltage reactions [20].

**9.2** Both electrode processes must be chemically and electrochemically compatible, i.e., no undesirable reactions should occur among the participating species [3].

**9.3** The selection of electrode materials is fundamental for the success of a paired system as it must take into account their possible interaction with the reagents, products and intermediates; a summary has been given elsewhere [2]. In addition, electrode materials must be selected carefully to prevent the global process from being controlled by one of the simultaneous reactions.

**9.4** For the selection of complementary reactions:

### 9.4.1 Concerning hydrogen production

#### 9.4.1a Promote hydrogen production

- When the aqueous chemistry of a paired cathodic reaction requires hydrogen atoms, it is suitable to use low hydrogen overpotential electrodes (e.g., Raney Ni). This also better current efficiency and lowers the necessary operating potential [57].

#### 9.4.1b Avoid hydrogen production

- In some cases, it is desirable to avoid hydrogen production by replacing it with a useful cathodic reaction to save assembly and consumption costs from

production to application, and to reduce the pressure in the electrolyzer and electricity consumption (e.g., nitrobenzene reduction) [24]. Another alternative consists of using cathodes with high overpotential for hydrogen generation.

- During the paired generation of dioxygen, the cathodic hydrogen production coupled to water oxidation can be replaced by the generation of  $\text{H}_2\text{O}_2$  and, therefore, the system is converted from a  $4e^-$  to a  $2e^-$  process. This translates into a considerable decomposition voltage decrease and up to a doubling of the production efficiency [36].  $\text{H}_2\text{O}_2$  can be significantly useful in other specific chemical transformations (see below).
- Hydrogen ions may be reduced simultaneously with the species of interest (e.g., nitrates) [39], thus competing for the active sites; this hinders their reduction. Alternatively, adsorbed hydrogen species may hydrogenate intermediates. Both cases are detrimental for the process of interest and this can be analyzed in advance by electrochemical techniques such as cyclic voltammetry [39].
- When using divided cells, the  $\text{H}^+$  generated during electrooxidation of organic compounds can be used in the cathodic chamber for electroreduction processes that require  $\text{H}^+$ . In this way, protons are not accumulated and hydrogen production is limited. A proton-exchange membrane can be used as separator to serve this objective [76].

### 9.4.2 Concerning dioxygen production

#### 9.4.2a Avoid dioxygen production

- The use of a hydrogen gas-fed anode has been reported to avoid the generation of dioxygen that consumes a large amount of energy [121].
- The use of acidic solutions is another strategy to inhibit dioxygen generation. This enhances current efficiency for target anodic processes [130].
- The use of high oxygen overvoltage electrodes (e.g., boron-doped diamond electrodes, BDDE) limits water oxidation and considerably expands the anodic electrochemical potential limit to allow the occurrence of other reactions of interest [103].

Once the paired reactions are selected, the necessary conditions for the individual reactions can be preliminarily optimized and the conditions for the paired process can be fully optimized afterwards [12].

## 10. Conclusions

Processes in which current is beneficially utilized at both electrode chambers in an electrochemical cell can result in significant advantages from the economic and environmental points of view. Since such processes are becoming more common,

design principles based on thermodynamic, kinetic, and chemical criteria -together with appropriate nomenclature- must be widely disseminated to facilitate their usage. Since the fraction of these processes that have reached industrial scale is minimal, a much greater emphasis needs to be given to such developments to benefit from the multiple advantages discussed in this paper. The electrochemical community is definitely developing more paired processes, which should be available in the near future. From the above discussion, it is clear that the organic synthesis field is reaching maturity while the inorganic synthesis and the environmentally related fields are emerging. The biochemical field is poorly exploited, and thus there is a vast opportunity in this area.

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