

Kinetic Spectrophotometric Method for the Estimation of Ofloxacin in Pharmaceutical Formulations

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Abstract. The objective of the current study was to develop a direct, sensitive spectrophotometric method based on the oxidation of Ofloxacin using potassium permanganate in alkaline medium. The rate of change of absorbance was measured at 603 nm. The initial rate method and fixed time method (at 4 min) are utilised to construct calibration graphs for calculating the concentration of the drug. The results were validated through inter day and intraday precision assays according to the ICH guidelines and also through recovery studies. Statistical comparison of the proposed methods with that of reference method shows excellent agreement and indicates no significant difference in their accuracy and precision.

Key words: Ofloxacin, Initial Rate Method, Fixed Time Method.

Resumen. Se desarrolla un método espectrofotométrico directo y sensible basado en la oxidación de ofloxacina con permanganato de potasio en medio alcalino. La velocidad de cambio de la absorbancia fue de 603 nm. Fueron utilizados los métodos de velocidad inicial y el método de tiempo fijo (4 min) para construir las gráficas de calibración para calcular la concentración del fármaco. Los resultados fueron validados inter e intra- días por ensayos de precisión según las pautas ICH y por estudios de recuperación. La comparación estadística de los métodos propuestos con el método de referencia muestra una excelente concordancia y no existe diferencia significativa de su exactitud y precisión.

Palabras clave: Ofloxacina, método de velocidad inicial, método de tiempo fijo.

Introduction

Ofloxacin is a synthetic fluoroquinolone antibiotic [1] and is chemically (*RS*)-7-fluoro-2-methyl-6-(4-methylpiperazin-1-yl)-10-oxo-4-oxa-1-azatricyclo [7.3.1.0⁵, 13] trideca-5(13),6,8,11-tetraene-11-carboxylic acid. It is prescribed for urinary tract infection, bronchitis, pneumonia, prostatitis, syphilis and infections of reproductive organs. Literature survey revealed the estimation of Ofloxacin has been determined along with other drugs by UV [2-4], HPLC [5-13], electrophoresis [14], HPTLC [15] and voltammetrically [16].

Kinetic spectrophotometric methods became of great interest in chemical and pharmaceutical analysis because of some specific advantages of the method, such as selectivity due to the measurements based on the absorbance with time of reaction, easily applicable, sensitive and the processing cost is low as the reagents used are commonly available in laboratories. There is no literature regarding estimation of ofloxacin using kinetic spectrophotometry. Therefore the objective of this study was to develop a sensitive method by employing the kinetic colorimetric oxidation of Ofloxacin to increase selectivity and consequently determination of low concentration of the drug as possible.

Materials and Methods

Apparatus

A Shimadzu UV-Visible spectrophotometer (model-1800, Japan) with a pair of matched quartz cells was used to measure absorbance.

Reagents and standards

All chemicals used were of analytical grade. Pure Ofloxacin was provided by Cipla, India and was used as received. A standard solution of Ofloxacin 1 mg/mL was prepared by dissolving 50 mg in 1 mL of 1 M NaOH and then making the volume with doubly distilled water up to 50 mL. A working test solution (0.1 mg/mL) of Ofloxacin was prepared by diluting 5 mL of the stock solution to 50 mL with doubly distilled water. The tablets containing Ofloxacin such as Oflox-200 (Cipla) and Oflocock-200 (Wockhardt) were purchased from local market. Aqueous solutions of sodium hydroxide (1M) and potassium permanganate (0.01M) were prepared with doubly distilled water. Potassium permanganate should be prepared freshly.

Procedures for the determination of ofloxacin

Initial rate method

Aliquots of 0.5-1.5 mL of 0.01% Ofloxacin test solution were pipette out into a series of flasks and to these 1 mL of 1 M NaOH and 1.5 mL of 0.01 M KMnO₄ was added and then diluted them with doubly distilled water up to 10 mL. The contents of the mixture were mixed well and their absorbance was measured at 603 nm with a function of time. The initial rate of reaction at different concentrations was obtained from the slope of tangent to the absorbance time curve. The calibration graph was plotted between the logarithm of the initial rate of reaction and the logarithm of the molar concentration of Ofloxacin. The amount of drug was calculated either from the calibration graph or regression equation.

Fixed time method

In this method, the absorbance of each sample solution at a preselected fixed time was accurately measured and plotted against the final concentrations of the drug. The content of the drug was calculated either from the calibration graph or regression equation.

Determination of Ofloxacin from dosage form

An accurately weighed quantity of tablets, equivalent to 50 mg of the drug dissolved in 1 mL of 1 M NaOH and volume made up to 50 mL of doubly distilled water and the residue was filtered using membrane filter. The stock solution was diluted according to need and analyzed by the recommended procedures.

Procedure for reference method

Into a series of 10 mL standard volumetric flask, different volumes (0.5-1.5 mL) of 0.01% drug (0.1 mg/mL) solution were pipette and diluted to volume with doubly distilled water. The absorbance was measured at absorption maximum of Ofloxacin that was found to be 289 nm. The amount of drug in a given sample was calculated from the calibration graph or regression equation. Recovery was performed for studying the accuracy of this method and the results are shown in Table 8.

Results and Discussion

Potassium permanganate acts as strong oxidising agent and has been used in oxidimetric analytical method for determination of many compounds. Basic principle involved during the due course of oxidation process is the alteration of valence of manganese. The heptavalent manganese ion (Mn VII) changes to green colour, MnO_4^{2-} (VI) in alkaline medium whereas in

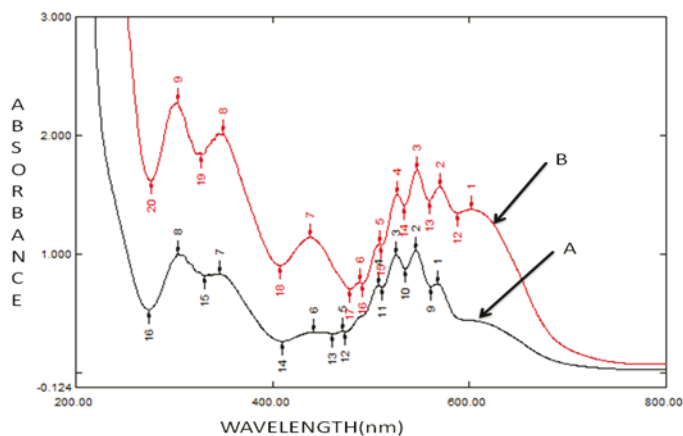


Fig. 1. Absorption spectra of (A) 1.0 mL of 1 M NaOH + 1.5 mL 0.01 M $KMnO_4$; (B) 15 $\mu g/mL$ of Ofloxacin + 1.0 mL of 1 M NaOH + 1.5 mL 0.01 M $KMnO_4$ in doubly distilled water.

neutral and acidic medium the permanganate is further reduced to colourless (Mn II). This behaviour of permanganate is the primary basis for its uses in development of spectrophotometric methods [17-18].

The absorption spectrum of aqueous potassium permanganate solution in the alkaline medium exhibited an absorption band peaking at 546 nm. The addition of Ofloxacin to this solution produced new characteristic band at 603 nm (Fig. 1). This band is primarily due to the formation of manganate ion, which resulted in the oxidation of Ofloxacin by potassium permanganate in alkaline medium. This could cause increase in intensity of the colour with time and this occurrence was utilized to develop a kinetics based analytical method for the determination of Ofloxacin. The different variables that affect the formation of manganate ion were studied and optimized.

Effect of $KMnO_4$ concentration

To study the effect of the $KMnO_4$ concentration, aliquots of Ofloxacin 100 $\mu g/mL$ were transferred into a series of 10 mL volumetric flasks, followed by varying volumes of 0.01 M $KMnO_4$ (0.2-2.0 mL) and 1 mL of 1 M NaOH solutions. The absorbance at 603 nm was measured at a fixed time of 16 min.

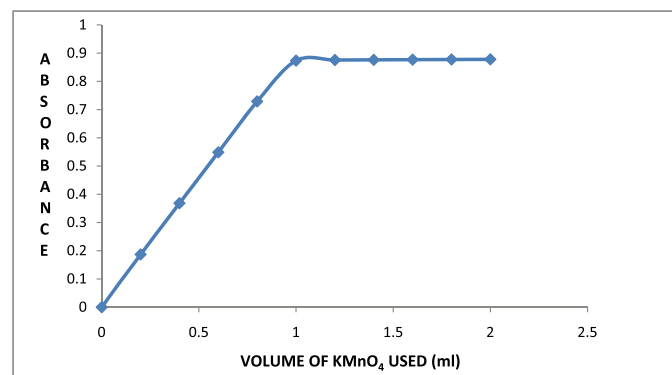


Fig. 2. Effect of the volume of 0.01 M $KMnO_4$ on the intensity of colour produced during the reaction (Ofloxacin 15 $\mu g/mL$; 1.0 mL of 1 M NaOH).

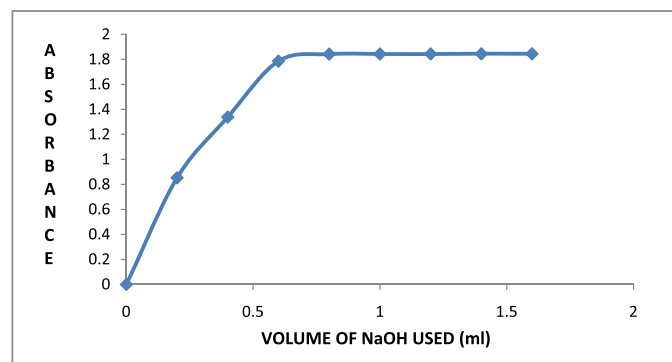


Fig. 3. Effect of the volume of 1 M NaOH on the intensity of colour produced during the reaction (Ofloxacin 15 $\mu g/mL$; 1.5 mL of 0.01 M $KMnO_4$).

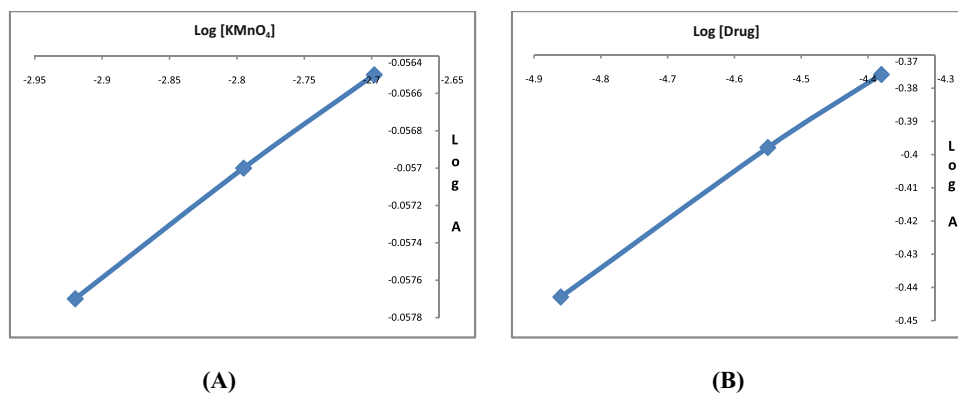
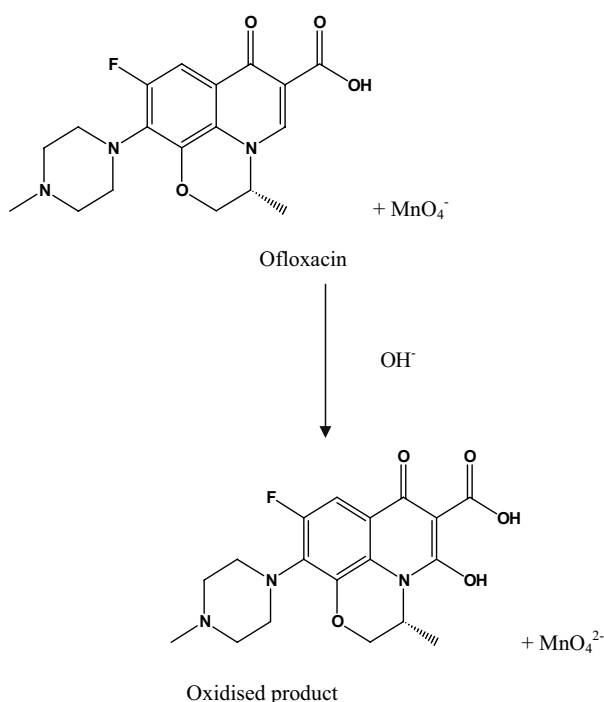


Fig. 4. Limiting logarithmic plot for molar combining ratio between Ofloxacin and KMnO₄: (A) log A vs. log [KMnO₄]; (B) log A vs. log [Drug].



Scheme 1.

It is apparent from Fig. 2 that the absorbance increased with increasing volume of KMnO_4 solution, and became more or less constant when the volume was 1.2 mL. Further, increase in the volume of KMnO_4 resulted in no change in the absorbance (Fig. 2). This results suggest that the adoption of 1.5 mL of KMnO_4 (0.01 M) in the final solution could be adequate for the highest concentration of Ofloxacin (15 $\mu\text{g/mL}$) used in its determination.

Effect of NaOH concentration

The influence of NaOH concentration on the formation of manganate ions was examined. aliquots of Ofloxacin 100 $\mu\text{g/mL}$ were transferred into a series of 10 mL volumetric flasks, followed by varying volumes of 0.01 M KMnO_4 (1.5 mL) and

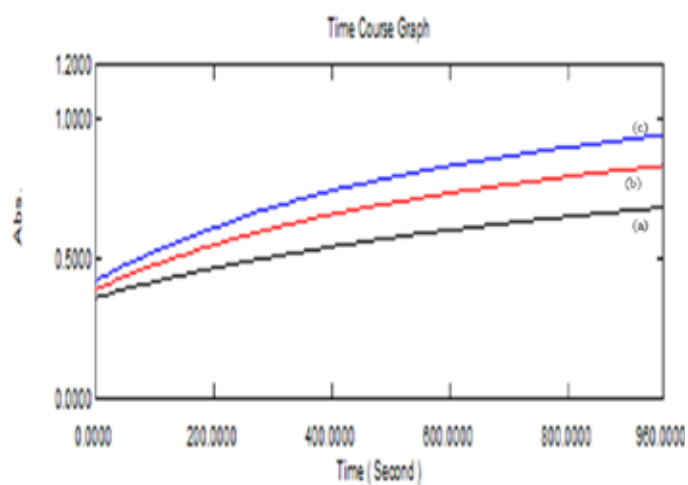


Fig. 5. Absorbance versus time graphs for the reaction of Ofloxacin with potassium permanganate. KMnO_4 is 1.5×10^{-3} mol/litre and Ofloxacin (a) 5 $\mu\text{g/mL}$ (b) 10 $\mu\text{g/mL}$ (c) 15 $\mu\text{g/mL}$ for 16 minutes.

(0.2-1.6 mL) of 1 M NaOH solutions. The absorbance at 603 nm was measured at a fixed time of 16 min. It is evident from the Fig. 3 that the highest absorbance of the reactant solution was obtained using 0.8 mL of 1M NaOH, in the current experimental condition. Hence, an optimum value of 1 mL of NaOH was selected and used in the further studies.

Effect of temperature

At room temperature the reaction rate increases substantially with time, although increasing the solution temperature enhanced the rate of reaction. However, we have observed that the increase in temperature led to increase in intensity of the colour.

Stoichiometry

The stoichiometry ratio between potassium permanganate and Ofloxacin was determined by the limiting logarithmic method. It was established by performing two types of experiments.

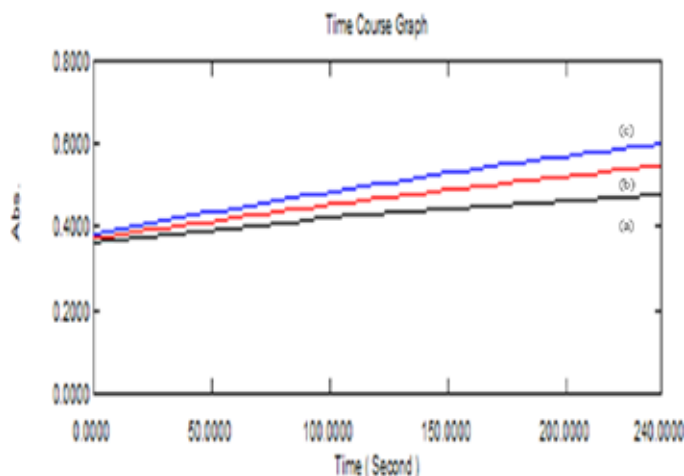


Fig. 6. Absorbance versus time graphs for the reaction of Ofloxacin with potassium permanganate. KMnO_4 is 1.5×10^{-3} mol/litre and Ofloxacin (a) 5 $\mu\text{g/mL}$ (b) 10 $\mu\text{g/mL}$ (c) 15 $\mu\text{g/mL}$ at a fixed time of 4 minutes.

In the first type, KMnO_4 concentration was kept constant and concentration of Ofloxacin was varied. In the second one KMnO_4 concentration was varied while keeping concentration of Ofloxacin, constant. The logarithm of the absorbance was plotted against the logarithm of the corresponding concentration. The slopes of the two straight lines were calculated and found to be unity in each case. Thus the stoichiometric ratio between ofloxacin and potassium permanganate was found to be 1:1 (Fig. 4).

Reaction

Literature revealed that fluoroquinolone antibiotics have a tendency to get oxidised by chloramine B, potassium permanganate in alkaline medium [19-20]. So, in this study we have used potassium permanganate (alkaline medium) as the oxidising agent for Ofloxacin. The reaction mechanism is given in Scheme 1.

Initial rate method

The initial rate of reaction was determined by measuring slopes of initial tangent to the absorbance time curves (Fig. 5). The order of reaction with respect to permanganate was determined from the reaction using different concentrations of KMnO_4 with fixed Ofloxacin concentration. The order of reaction was found to be one, as the plot of initial rate against the initial absorbance was linear and was passing through the origin. Similarly, the order of Ofloxacin was determined by plotting the logarithm of the initial rate of the reaction with the logarithm of the molar concentration of Ofloxacin. Here also the order was found to be one.

Under the optimized experimental conditions, the concentration of Ofloxacin was determined using an excess of KMnO_4 and NaOH solution with respect to the initial concentration

Table 1. Optical characteristics of fixed time method.

Parameters	Fixed time method			
	4 min	8 min	12 min	16 min
Beer's law limit ($\mu\text{g/mL}$)	3-15	3-9	3-9	3-9
Regression equation	$Y = mX + c$			
Slope	0.373	0.594	0.584	0.643
Intercept	0.299	0.295	0.364	0.395
Correlation coefficient (r)	0.999	0.996	0.999	0.998
Standard deviation	0.050	0.058	0.081	0.093
Relative standard deviation	0.079	0.089	0.114	0.119
LOD ($\mu\text{g/mL}$)	0.447	0.322	0.457	0.477
LOQ ($\mu\text{g/mL}$)	1.356	0.976	1.386	1.446

of Ofloxacin, under the optimized experimental condition. As a result, a pseudo zero order reaction with respect to their concentrations. However, the initial rate of the reaction would follow a pseudo first order and was found to obey the following equation;

$$\text{Rate} = \Delta A / \Delta t = K^1 C^n$$

Where K^1 is the pseudo first order rate constant, C is the concentration of Ofloxacin, n is the order of the reaction. The logarithmic form of the above equation is written as,

$$\text{Log (rate)} = \text{log } K^1 + n \text{ log } C$$

The linear regression analysis using the method of least square was used to determine the slope, intercept and correlation coefficient. The calibration curves plotted between the log of initial rate versus log of molar concentration of Ofloxacin were linear over a range of 5-15 $\mu\text{g/mL}$ with a correlation coefficient of 0.999.

Fixed time method

In this method, the absorbance of coloured solutions with different amount of Ofloxacin was measured at a preselected time (Fig. 6). Calibration curve of absorbance versus initial concentration of Ofloxacin were plotted at a fixed time of 4, 8, 12 and 16 minutes. The regression equations, correlation coefficient, limit of detection, standard deviation and relative standard deviation are summarized in Table 1.

The lowest detection limit was obtained at 8 min where as the widest concentration range of quantification was obtained at a fixed time of 4 min. According to ICH guidelines [21] the detection limit is not required to be the part of the validation procedure for assays. Therefore on the basis of wider concentration range and less time of analysis, the fixed time of 4 min was selected for analysis.

The accuracy and precision of Ofloxacin was measured from the pure sample and tablet dosage form three times within

Table 2. Intra-day assay for the evaluation of accuracy and precision of the initial rate and fixed time method.

Methods	Amount taken ($\mu\text{g/mL}$)	Amount found ($\mu\text{g/mL}$)	% Accuracy	% RSD
Initial rate method	5	4.922	98.44 ± 0.939	0.954
	10	10.207	102.07 ± 0.655	0.642
	15	14.869	99.12 ± 0.144	0.145
Fixed time method	5	5.05	101 ± 1.019	1.009
	10	10.22	101.2 ± 0.588	0.581
	15	15.14	100.95 ± 0.874	0.866

Values are mean \pm SD for 3 determinations.

Table 3. Intra-day assay for the test of precision of proposed method in pharmaceutical formulations.

Formulations	Initial rate method				Fixed time method			
	Amount taken ($\mu\text{g/mL}$)	Amount found ($\mu\text{g/mL}$)	% Accuracy \pm SD	% RSD	Amount taken ($\mu\text{g/mL}$)	Amount found ($\mu\text{g/mL}$)	% Accuracy \pm SD	% RSD
Oflox-200	5	4.994	99.88 ± 0.226	0.226	5	5.02	100.53 ± 2.07	2.059
	10	10.262	102.59 ± 0.641	0.620	10	10.17	101.7 ± 2.67	2.625
	15	15.11	100.73 ± 0.424	0.420	15	15.08	100.53 ± 1.256	1.249
Oflocock-200	5	4.937	98.74 ± 0.742	0.751	5	5.04	100.86 ± 0.805	0.798
	10	10.31	103.18 ± 0.605	0.586	10	9.97	99.73 ± 1.020	1.023
	15	14.879	99.19 ± 0.898	0.905	15	15.1	100.70 ± 0.491	0.488

Values are mean \pm SD for 3 determinations.

Table 4. Inter- day assay for the evaluation of accuracy and precision of the initial rate and fixed time method.

Methods	Amount taken ($\mu\text{g/mL}$)	Amount found ($\mu\text{g/mL}$)	% Accuracy \pm SD	% RSD
Initial rate method	5	4.982	99.77 ± 0.121	0.121
	10	10.27	102.67 ± 0.767	0.747
	15	15.005	100.04 ± 0.290	0.289
Fixed time method	5	5.11	102.26 ± 1.319	1.290
	10	10.06	100.62 ± 1.305	1.297
	15	15.27	101.82 ± 1.312	1.289

Values are mean \pm SD for 3 determinations.

Table 5. Inter-day assay for the test of precision of proposed method in pharmaceutical formulations.

Formulations	Initial rate method				Fixed time method			
	Amount taken ($\mu\text{g/mL}$)	Amount found ($\mu\text{g/mL}$)	% Accuracy \pm SD	% RSD	Amount taken ($\mu\text{g/mL}$)	Amount found ($\mu\text{g/mL}$)	% Accuracy \pm SD	% RSD
Oflox-200	5	4.982	99.64 ± 0.103	0.103	5	5.04	100.8 ± 1.697	1.684
	10	10.275	102.75 ± 0.470	0.457	10	10.21	102.1 ± 1.433	1.404
	15	14.995	99.96 ± 0.383	0.383	15	15.17	101.13 ± 1.225	1.211
Oflocock-200	5	4.938	98.76 ± 0.942	0.953	5	5.06	101.26 ± 0.659	0.651
	10	10.343	103.43 ± 0.801	0.774	10	10.06	100.63 ± 0.928	0.922
	15	14.95	99.66 ± 1.084	1.087	15	15.17	101.15 ± 1.092	1.080

Values are mean \pm SD for 3 determinations.

Table 6. Recovery studies for the determination of Ofloxacin in tablet dosage form by initial rate method.

Formulations	Initial rate method				
	Amount taken ($\mu\text{g/mL}$)	Amount added ($\mu\text{g/mL}$)	Amount found ($\mu\text{g/mL}$) \pm SD	% Recovery	% RSD
Oflox -200	5	4	9.01 \pm 0.018	100.46	0.200
		5	9.98 \pm 0.057	99.68	0.571
		6	11.08 \pm 0.016	101.38	0.144
Oflocock-200	5	4	9.03 \pm 0.016	100.75	0.177
		5	10.02 \pm 0.051	100.57	0.509
		6	11.09 \pm 0.004	101.55	0.036

Values are mean \pm SD for 3 determinations.

Table 7. Recovery studies for the determination of Ofloxacin in tablet dosage form by fixed time method.

Formulations	Fixed time method				
	Amount taken ($\mu\text{g/mL}$)	Amount added ($\mu\text{g/mL}$)	Amount found ($\mu\text{g/mL}$) \pm SD	% Recovery	% RSD
Oflox-200	5	4	8.99 \pm 0.002	99.29	0.022
		5	10.02 \pm 0.012	100.06	0.12
		6	10.99 \pm 0.0008	99.56	0.007
Oflocock-200	5	4	9.07 \pm 0.055	100.41	0.606
		5	10.06 \pm 0.041	100.06	0.408
		6	11.12 \pm 0.081	101.05	0.728

Values are mean \pm SD for 3 determinations.

Table 8. Recovery studies for the determination of Ofloxacin using reference method.

Formulations	Reference method				
	Amount taken ($\mu\text{g/mL}$)	Amount added ($\mu\text{g/mL}$)	Amount found ($\mu\text{g/mL}$) \pm SD	% Recovery	% RSD
Oflox-200	5	4	8.98 \pm 0.040	100	0.445
		5	10.06 \pm 0.047	100.6	0.467
		6	10.98 \pm 0.044	100.16	0.401
Oflocock-200	5	4	9.06 \pm 0.040	100.25	0.442
		5	10.01 \pm 0.032	99.2	0.320
		6	11.07 \pm 0.042	100.33	0.379

Values are mean \pm SD for 3 determinations.

Table 9. Point hypothesis: comparison of the proposed method with reference method using paired student *t*-test.

Formulations	Initial rate method			Fixed time method			Reference method	
	% Recovery	% RSD	t-value	% Recovery	% RSD	t-value	% Recovery	% RSD
Oflox-200	100.46	0.200	0.420	99.29	0.022	0.16	100	0.445
	99.68	0.571	0.370	100.06	0.12	0.42	100.6	0.467
	101.38	0.144	0.090	99.56	0.007	0.39	100.16	0.401
Oflocock-200	100.75	0.177	0.980	100.41	0.606	0.79	100.25	0.442
	100.57	0.509	0.670	100.06	0.408	0.88	99.2	0.320
	101.55	0.036	0.105	101.05	0.728	0.75	100.33	0.379

Degree of freedom = 2

Theoretical t-value (n = 3) at 95% confidence level is 1.943.

a day by initial rate and fixed time method are given in Table 2 and 3 respectively. The daily precision was measured by analysing the pure sample and tablets on three consecutive days by both methods (Table 4 and 5 respectively). The standard deviation, relative standard deviation obtained in both intraday and inter-day assays was found to be satisfactory.

The accuracy of the methods was determined by performing recovery studies. For this a known amount of pure drug was added to the preanalysed tablet dosage form and then reanalysed by the recommended procedure. The results obtained are shown in Table 6 and 7. However, there was no interference by the formulations excipients.

Application of the proposed method to pharmaceutical dosage form

The initial rate and fixed time methods of the proposed kinetic spectrophotometric method for the determination of ofloxacin have been tested on commercial pharmaceutical dosage forms (tablets) and concentration of ofloxacin was computed from its corresponding regression equation. The results of the proposed method were then compared with the reference method using paired student t-test. The values of the t-test obtained are given in Table 9 which were less than the theoretical values and thus confirms that there was no difference in the performance of the proposed method and the reference method.

Advantages of the proposed method

The proposed method involves measurements in the visible region and is more selective than the reported UV based methods. Further, the proposed method utilizes water as solvent which is easily available, safe and economical as well [20]. In contrast, the other methods reported used costly organic solvents like methanol, ethanol, which cause health and environmental hazards and they require special care for their sanitary disposal. Moreover, the current method is very simple as it is a single step reaction and is not tedious.

From economical point of view, all the analytical reagents are inexpensive, have excellent shelf life, and are available in any analytical laboratory. The process used in the proposed method need not require any tedious liquid-liquid extraction for the drug base or extraction of the chromophores [22].

Conclusion

The initial rate and fixed time method can be easily applied to Ofloxacin in its determination in pure and tablet dosage form. These methods do not require any elaborate treatment of the analyte. The proposed methods are sensitive enough for the determination of lowest amount of drug. These advantages

encourage the application of the proposed method in routine analysis of Ofloxacin in laboratories.

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