



Original

Kinetic studies of surface modification of lignocellulosic *Delonix regia* pods as sorbent for crude oil spill in water

Jude Chinedu Onwuka^{a,*}, Edith Bolanle Agbaji^b, Victor Olatunji Ajibola^b, Friday Godwin Okibe^b

^a Department of Chemistry, Federal University Lafia, Nasarawa, Nigeria

^b Department of Chemistry, Ahmadu Bello University Zaria, Kaduna, Nigeria

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Abstract

Acetylation method was used in the modification of *Delonix regia* pods (DRPs) as sorbent for crude oil cleanup in water. Fourier transform infra-red (FTIR) and scanning electron microscope (SEM) analysis were used to investigate the influence of acetylation and crude oil sorption on the sorbent (DRPs). Reaction conditions played significant roles on the extent of acetylation of DRPs. Temperatures of 303 K and 343 K were found to be the most suitable acetylation temperatures of DRPs. Intra-particle diffusion was the rate controlling mechanism for acetylating DRPs while the contributing mechanisms depend on the temperature of acetylation. The crude oil sorption capacity (OSC) of modified DRPs was significantly higher than that of unmodified DRPs. Physical and chemical reactions were faster in the crude oil sorption by modified DRPs than the unmodified while diffusion into the pores of the modified DRPs was slower than in the unmodified. Hydrophobic functional groups were enhanced by acetylation and crude oil molecules were adsorbed at these functional groups. Surface structure, pore sizes and fiber lengths of the sorbent were affected by acetylation and crude oil sorption. FTIR and SEM showed clear evidence of successful acetylation and crude oil sorption. Analysis of variance (ANOVA) was used to determine the statistical difference of weight percent gain (WPG) obtained from the acetylation of DRPs at different reaction parameters such as temperature and time. The OSC of modified and unmodified sorbents at various contact times, were also compared using ANOVA.

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1. Introduction

Nigeria economy is dependent on crude oil exploration at the Niger delta region of the country. Recent agitations for resource control and sovereign state by the people of this region, has led to the bombing of oil installations (especially crude oil pipelines) by emerging militants groups in this region. This, in addition to other accidental and intentional crude oil discharges into aquatic the environment, result to severe negative impact on the aquatic environment due to the toxic effect many compounds in crude oil, poses to aquatic organisms (Aguilera, Méndez, Pásaroa, & Laffona, 2010; Alonso-Alvarez, Pérez, & Velando, 2007). Thus,

oil spill in water bodies remains a serious concern to both local and international communities. Regulations (US EPA, 2008) have done much to prevent oil waste contamination during transport on the open seas. When crude oil spill in water surface does occur, the issue is not only the cleaning of the aquatic environment but also recovery of this precious commodity. Hence, any oil absorbing material used must also be able to release the oil. Thus, it is necessary to develop a relatively efficient method to eliminate the potential hazards of crude oil spill in water and recover the spilled crude oil simultaneously. Until now, oil spills skimmers, oil dispersant, oil gelling agent and oil absorbent have been used for crude oil spill clean-up in water (Wang, Zheng, & Wang, 2012).

The effectiveness of the sorbents in saturated aquatic environments would be enhanced if the density of the hydroxyl functional groups is decreased (Bodirlau & Teaca, 2009). This functional group (hydroxyls) is the most reactive and abundant

* Corresponding author.

E-mail address: emperor20062003@yahoo.com (J.C. Onwuka).

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site in the cell wall polymers of the lignocellulosic materials. The hydroxyl functionality of these fibers can be reduced by chemical modification such as acetylation, methylation, cyanoethylation, benzoylation, acrylation, acylation, etc. (Breitenbeck, Grace, Holiday, & Assoc. 1997; Sun, Sun, & Sun, 2004). According to Diya'uddeen, Mohammed, Ahmed, and Jibril (2008), sorbents with large surface areas and affinity to organic compound could be developed from cost effective and readily available agricultural by-products.

Huge amounts of agricultural wastes are produced in Nigeria. However, only a fraction of these materials are 'usefully' reused because they are thereafter used majorly as domestic sources of fuel (cooking) (Nwadiogbu, Okoye, Ajiwe, & Nnaji, 2014).

Delonix regia with common names; Pride of Barbados, dwarf Poinciana, Bird of Paradise, flamboyant-de-jardin (Prohp et al., 2004) and *waken bature* by the Hausas, is abundant in Nigeria. It is a legume belonging to the Leguminosae family, which is the second largest family among the dicotyledonous plants (Prohp et al., 2004). The bark of *D. regia*, has medicinal properties while the hard, elongated seeds are occasionally used as beads (Okenwa & Louis, 2014). The plant produces brown woody seed pods purely a waste material (Yuh-Shan & Malarvizhi, 2009).

This study therefore investigates the kinetic pathway to modification of *D. regia* pods by acetylation and the efficiency of the modified sorbent (DRPs) in the removal of crude oil from water surface.

2. Methodology

2.1. Sample collection and preparation

D. regia pods (DRPs) were collected from the environment of National Research Institute for Chemical Technology (NAR-ICT) Zaria, Kaduna – Nigeria. It was cut and ground in a mortar. Then, thoroughly washed with distilled water to remove foreign materials, and water soluble components. The washed DRPs was initially air dried for 24 h and later oven dried to a constant weight, at 338 K for 48 h.

After drying, it was sieved to obtain a homogenous particle size (i.e. 425–625 μm) using the BS410/1986 laboratory test sieve. The sieve was shaken using a mechanical shaker.

2.2. Acetylation of the sorbent (DRPs)

A portion (0.5 g) of sorbent (DRPs) was placed in a 50 mL conical flask containing 20 mL of acetic anhydride and 0.2 g (1% of the solvent) N-bromosuccinimide (NBS). Effect of catalyst (NBS) on acetylation of the sorbent was done using different quantities of catalyst (0, 1, 2, 3 and 4% of acetic anhydride used) at constant sorbent dose, time, temperature and pressure. Effect on varying concentration of the sorbent (0.1, 0.25, 0.50, 0.75 and 1.0 g) was studied at constant amount of catalyst, time, temperature and pressure. Effect of reaction time was carried out by acetylating the sorbent at various reaction times (30, 60, 90, 120 and 150 min). Time effect was studied under various temperatures (303, 323, 343 and 363 K) while all other factors were kept constant in a temperature controlled water bath. After

each batch experiment, the conical flask was removed from the water bath and the hot filtrate was decanted off. The residue (sorbent) was thoroughly washed with ethanol and acetone to remove unreacted acetic anhydride and acetic acid by-product. The products were oven dried at 333 K for 16 h, and later cooled in a desiccator and stored in a plastic container prior to analysis. The extent or level of modification of the sorbent due to acetylation was estimated using weight percent gain (WPG).

2.2.1. Weight percent gain

The weight percent gain (WPG) was determined by gravimetric method as described by Thompson, Emmanuel, Adagadzu, and Yusuf (2010); Azeh, Olatunji, Mohammed, and Mamza (2013). It was calculated on the basis of oven-dried unreacted sorbent (DRPs) fibers. The dried sorbent obtained were reweighed to determine the weight gain on the basis of initial oven dry measurements. Weight percent gain (WPG) of the sorbent due to acetylation was calculated using the expression

$$\text{WPG (\%)} = \left[\frac{\text{Weight gain}}{\text{Original weight}} \right] \times 100 \quad (1)$$

2.3. Oil sorption capacity

To simulate the situation of oil spill and minimize experimental variation, the crude oil sample was held in beakers for one day in open air to release volatile hydrocarbon contents. The raw and acetylated samples were subjected to crude oil sorption test.

To 100 mL of distilled water in a 250 mL beaker, 2.5 g of crude oil was added. A portion (0.5 g) of the sorbent was added into the mixture in the beaker and left unperturbed for 10 min. After 10 min, the sorbent was removed using sieving net and left to drain by hanging the net over the beaker in an oven for 4 h at 333 K. The drained sorbent was weighed and recorded. This was repeated at different times (5, 10, 15, 20 and 25 min) at constant concentration and also at different initial concentrations of crude oil (1.25, 2.5, 3.75, 5.0 and 6.75 g/100 mL of water) at constant time. The sorption capacity of the sorbent samples was calculated using the expression

$$\text{Oil sorption capacity} = \frac{\text{New weight gained}}{\text{Original weight}} \quad (2)$$

and was recorded as gram per gram of sorbent. The procedure was carried out in triplicates and the mean of the results reported.

2.4. Analysis of variance (ANOVA)

ANOVA test compares two variance estimations: variance within group (the unsystematic variation or error in the data) and variance between groups (effects due to the experiment) (Hinkle, Wiersma, & Jurs, 1998).

In this study, independent variable (acetylation reaction time and oil sorption contact time, variation) and dependent variable (weight percent gain and oil sorption capacity) were compared using ANOVA.

2.5. Scanning electron microscope (SEM) analysis

A Phenom World model of SEM machine was used in this analysis. About 25 mg of the oven dried sample was sputter coated with a gold layer in a sputter machine (quantum sputter). The SEM machine was allowed to stabilize for 120 s before setting the parameters to be used. Imaging the sample was done at 15 kV set at 1000 magnification.

2.6. Fourier transform infra-red (FTIR) analysis

The translucent sample disk used in the FTIR spectra analysis was prepared by encapsulating 30 mg of finely ground sorbent in 300 mg of KBr. The FTIR spectra was recorded using Shimadzu-8400S Fourier transform infrared spectrometer (FT-IR) over the spectra range of 4000–500 cm^{-1} with a resolution of 4 cm^{-1} . This was carried out at the National Research Institute for Chemical Technology (NARICT) Zaria.

3. Theoretical consideration of reaction kinetics

In order to investigate the mechanism of acetylation, crude oil sorption process and the potential rate controlling steps, such as mass transport, pore diffusion and chemical reaction processes, kinetic models have been used to fit experimental data.

3.1. Pseudo-first order kinetics

On the assumption that the acetic anhydride reagent is present in large quantities for surface –OH sites, rate of reaction becomes dependent upon the concentration of –OH groups (Hill, Jones, Strickland, & Cetin, 1998), which is proportional to the WPG due to acetylation. Thus, pseudo-first order equation considered for acetylation process can be expressed as

$$\ln \text{WPG}_t = \ln \text{WPG}_0 - k_1 t \quad (3)$$

Similarly, for crude oil sorption kinetics, Eq. (3) is represented as

$$\ln \text{OSC}_t = \ln \text{OSC}_0 - k_1 t \quad (4)$$

where WPG_0 is the minimum weight percent gain, WPG_t is the weight percent gain at time t , OSC_0 is the minimum oil sorption capacity of the sorbent, OSC_t is the oil sorption capacity of the sorbent at time t , t is the time in minutes, k_1 is pseudo first order rate constant. The plot of $\ln \text{WPG}_t$ against t gives a linear relationship representing Eq. (3), where $-k_1$ is the slope and $\ln \text{WPG}_0$ is the intercept. Using Eq. (4), similar plot can be made for crude oil sorption kinetics. If the coefficient of regression obtained in this plot is moderate (i.e. $0.43 \leq R^2 \leq 0.83$), it implies that the mechanism at that particular temperature is due to surface reaction or physisorption (Nwadiogbu et al., 2014).

3.2. Pseudo-second order kinetics

The required pseudo-second order equation for acetylation is given as

$$\left[\frac{1}{\text{WPG}_t} \right] = k_2 t + \left[\frac{1}{\text{WPG}_0} \right] \quad (5)$$

For crude oil sorption, Eq. (5) becomes

$$\left[\frac{1}{\text{OSC}_t} \right] = k_2 t + \left[\frac{1}{\text{OSC}_0} \right] \quad (6)$$

The plot of $(1/\text{WPG}_t)$ against t in Eq. (5), gives a linear relationship with k_2 as the slope and $(1/\text{WPG}_0)$ is the intercept. Eq. (6) gives similar plot for crude oil sorption kinetics. If the coefficient of regression obtained in this plot is moderate, it implies that chemisorption (Chemical reaction) contributes or is the rate controlling step for the mechanism at that particular time or temperature as the case may be.

3.3. Intra-particle diffusion

The possibility of intra-particle diffusion being the rate determining step are explored using the intra-particle diffusion model (Srivastava, Swamy, Mall, Prasad, & Mishra, 2006). This involves the transport of the adsorbate to the surface of the adsorbent particles and the diffusion of the solute molecules into the interior of the pores, which is usually a slow process. The kinetic results are analyzed by the intra-particle diffusion model to elucidate the diffusion mechanism. Equation representing this model is expressed as: (Igwe & Abia, 2006; Subbaiah, Kalyani, Reddy, Boddu, & Krishnaiah, 2008)

$$\text{WPG}_t = k_3 \sqrt{t} + c \quad (7)$$

For crude oil sorption, Eq. (7) becomes

$$\text{OSC}_t = k_3 \sqrt{t} + c \quad (8)$$

where k_3 is the intra-particle diffusion rate constant, which can be evaluated from the slope of the linear plot of WPG_t versus $t^{1/2}$ and c is the intercept. Crude oil sorption kinetics can be modeled with similar plot using Eq. (8). The intercept of the plot reflects the boundary layer effect. The larger the intercept, the greater the contribution of the surface sorption in the rate controlling step (Eba et al., 2010). If the regression of WPG_t versus $t^{1/2}$ is linear and passes through the origin, then intra-particle diffusion is the sole rate-limiting step. However, in the case where the line do not pass through the origin, it is suggested that the intra-particle diffusion is not the only mechanism involved in the sorption process due to some degree of boundary layer control (Bulut, Ozacar, & Sengil, 2008).

3.4. Liquid film-diffusion

When there is transport of the sorbate molecules from the liquid phase to the solid phase, boundary plays a major role in

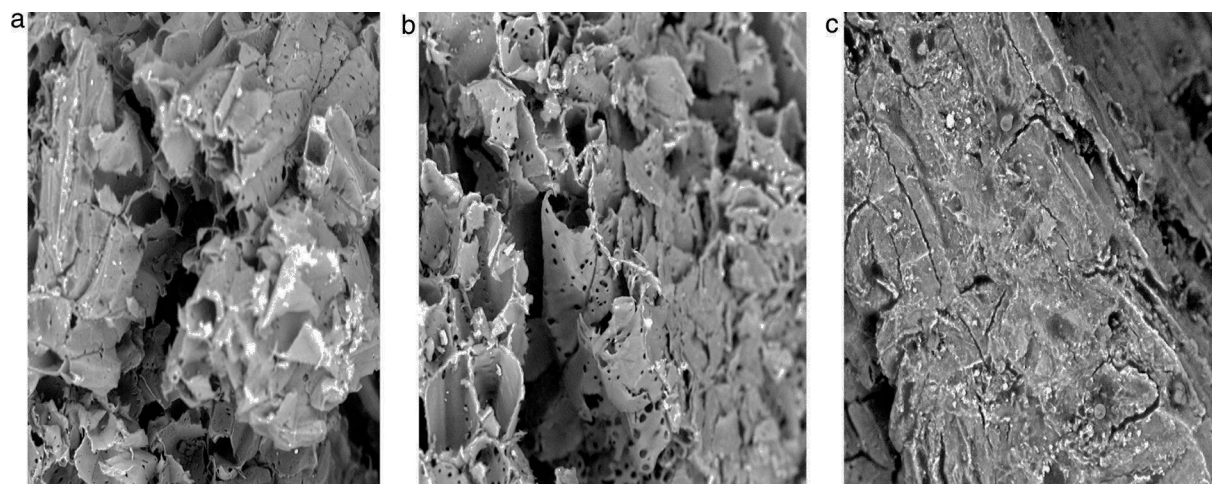


Fig. 1. SEM micrograph for unmodified (a), modified (b) and crude oil treated (c) DRPs.

adsorption, the liquid film diffusion model may be applied as follows: (Bulut et al., 2008; Eba et al., 2010)

$$-\ln(1 - F) = k_4 t \quad (9)$$

where k_4 is the liquid film-diffusion rate constant, F is fractional attainment which is equal to WPG_t/WPG_e for acetylation process and OSC_t/OSC_e for crude oil sorption process; WPG_t is the WPG due to acetylation at a particular time, WPG_e is WPG due to acetylation at equilibrium, OSC_t is the oil sorption capacity of the sorbent at a particular time and OSC_e is oil sorption capacity of the sorbent at equilibrium. A linear plot of $-\ln(1 - F)$ against t with zero intercept would suggest that the kinetics of the sorption process is controlled by diffusion through liquid film surrounding the solid sorbent. The small intercepts of liquid film-diffusion plot will suggest that liquid film-diffusion model might have some roles to play in the kinetics of acetylation (Eba et al., 2010).

4. Results and discussion

4.1. Scanning electron microscope (SEM) analysis

Comparing Figure 1a–c, it was observed that the surface morphology of the sorbent (DRPs) changed after different treatment processes. Figure 1b showed that the surface of modified sorbent was more ruptured along with different degree of wrinkles and grooves which increases the surface area. There are noticeable increase in the number pores in the sorbent (Fig. 1b) after acetylation. It was also observed from Figure 1c that the hollow lumen which can store crude oil reduced significantly. In other words, the surface of crude oil treated sorbent are smoother than that of unmodified and modified sorbents which further proves adsorption of crude oil on these surfaces. Wang et al. (2012) reported similar findings.

4.1.1. Pore analysis

Figure 2a–c showed that the various pores in the sorbents (DRPs) were grouped into three different size ranges. The histogram shown in these Figures showed that the sorbents are

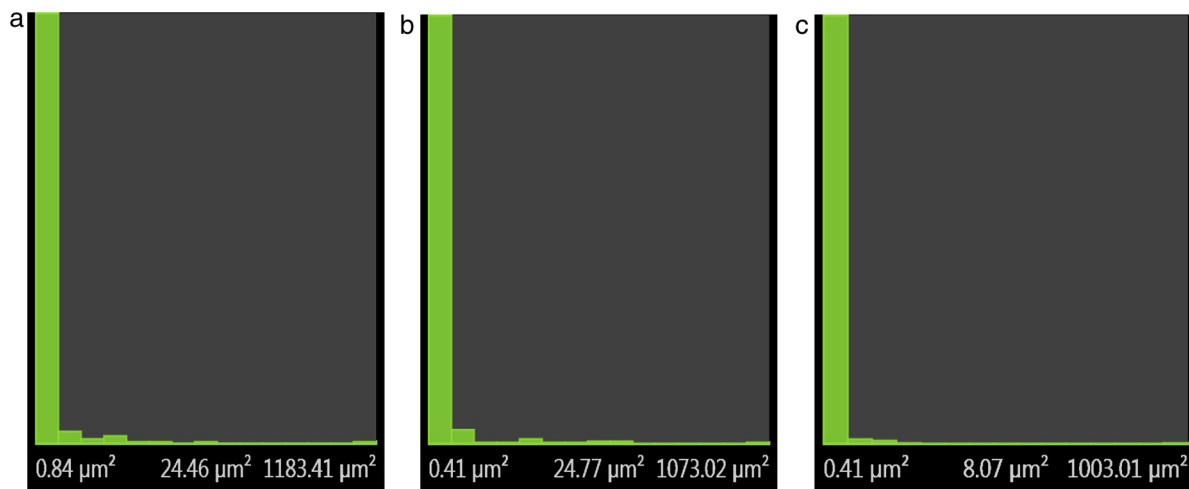


Fig. 2. Pore histogram for unmodified (a), modified (b) and crude oil treated (c) DRPs.

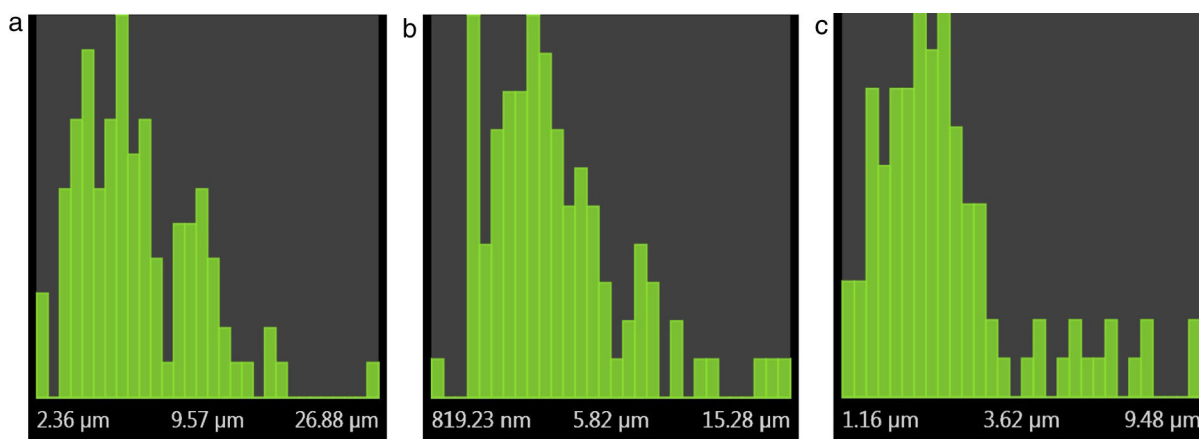


Fig. 3. Fiber histogram for unmodified (a), modified (b) and crude oil treated (c) DRPs.

composed of mainly small pores while medium and large pores are present in negligible amounts compared to the small pores. It was observed that acetylation significantly reduces the size of small and large pores, from 0.84 and 1183.41 μm^2 to 0.41 and 1073.02 μm^2 respectively while it increases the size of the medium pores. The reduction in the size of the small and large pores could be attributed to the expansion (or opening up) of the medium pores by acetylation. This shows that acetylation opens up the medium pores of the sorbent and thus, increases the surface area of the sorbent which subsequently enhances the crude oil sorption capacity of the sorbent. However, since the small pores are highly present in the sorbents, reduction in its size and that of large pores will logically lead to reduction in the overall porosity of the sorbent.

Figure 2c showed that the small pores were unaffected by crude oil sorption. However, crude oil sorption reduced medium and large pores significantly. Thus, crude oil was sorbed significantly at the medium pores.

4.1.2. Fibermetrics

Figure 3a–c showed that the fibers are grouped into three different lengths. It showed that the lengths of the different groups of fibers reduced after acetylation. This could be attributed to the swelling of the fibers as a result of attachment of high

molecular acetyl groups. This led to increase in the fiber length which subsequently breaks into short length fibers. Thus, increasing the amount of the short and medium length fibers after acetylation as shown in Figure 3b. Figure 3c showed that crude oil sorption resulted to the production of more short length fibers compared to the medium and long fibers. This could be attributed to high swelling impact large molecules of crude oil causes on the medium and long fibers. The size of the short fibers was observed to increase after crude oil sorption due to the swelling of the fibers.

4.2. Fourier transform infra-red (FTIR) spectra analysis

Figure 4 represents the IR spectra of unmodified, modified and crude oil treated sorbent (DRPs). The assignment of the observed signals or band are presented in Table 1.

FTIR spectra of the unmodified and modified sorbent showed evidence of acetylation with three intense ester bands appearing and/or enhanced at 1737–1751 cm^{-1} (carbonyl C=O stretching of ester), 1371–1388 cm^{-1} (C–H in $-\text{O}(\text{C}=\text{O})-\text{CH}_3$) and 1238–1255 cm^{-1} (C–O stretching of acetyl group) (Adebajo & Frost, 2004; Nwadiogbu et al., 2014). The absence of absorption peak in the region 1840–1760 cm^{-1} in the spectra of the

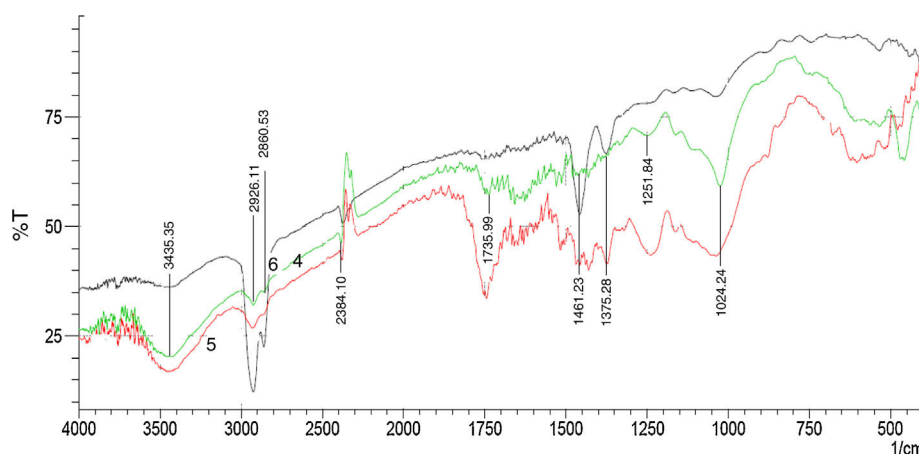


Fig. 4. FTIR spectra for unmodified [green (4) spectra] and modified [red (5) spectra] and crude oil treated [black (6) spectra] DRPs.

Table 1
Assignment of IR bands of functional groups in the sorbent.

Band	Sample 4 (cm ⁻¹)	Sample 5 (cm ⁻¹)	Sample 6 (cm ⁻¹)	Assignment
1	1024.24	1037.74	1039.67	C–O stretching vibrations in cellulose, hemicellulose and primary alcohol.
2	1109.11	–	–	C–O stretching from carboxylate groups of cellulose and hemicellulose
3	1161.19	1161.19	–	
3	1251.84	1238.34	1251.84	C=O stretch, acetyl group (lignin)
4	–	1321.28	–	
5	–	1371.43	1375.29	C–H deformation in –O–C=O–CH ₃ group.
6	1431.23	1431.23	1458.23	
7	–	1469.81	–	
8	1529.6	1518.03	–	–C=O stretching vibrations of aromatic rings caused by lignin
10	1656.91	1647.26	–	H–O–H bending of absorbed water
11	1735.99	1745.64	1741.78	C=O of esters (ascribed to lignin and hemicellulose)
13	2278.01	–	–	CH ₃ group, stretching vibrations of aliphatic CH ₃ group
15	2384.1	2378.31	2378.31	
16	2860.53	2864.39	2860.53	C–H stretch (methyl and methylene groups cellulose)
17	2926.11	2931.9	2926.11	
18	3435.34	3444.98	3431.48	O–H stretching vibrations in cellulose and hemicelluloses.
19	–	–	–	
20	3767.1	–	–	

Key: Samples 4, 5 and 6 represents unmodified, modified and crude oil treated DRPs respectively.

Ref: Nwadiogbu et al. (2014), Kudaybergenov et al. (2012), Azeh et al. (2013), Chung et al. (2011), Thompson et al. (2010), Adebajo & Frost (2004).

modified sorbent indicates that the modified sorbent is free of unreacted acetic anhydride.

The spectra of crude oil treated sorbent show the enhancement of sharp peaks at about 1375.29–1458.23 cm⁻¹ (C–H in –O(C=O)–CH₃), 2278.01–2384.10 cm⁻¹ (CH₃ group, stretching vibrations of aliphatic CH₃ group) and 2860.53–2926.11 cm⁻¹ (C–H in methyl and methylene groups) which confirms that crude oil is absorbed at the hydrophobic groups of the modified sorbent (Ibrahim, Ha-Ming, & Wang, 2009; Kudaybergenov, Ongarbayev, & Mansurov, 2012). Thus, confirming the need for modification by acetylation. This is consistent with the FTIR spectra of acetylated cellulosic materials reported by other researchers (Adebajo & Frost, 2004; Azeh et al., 2013; Chung, Suidan, & Venosa, 2011; Nwadiogbu et al., 2014).

4.3. Acetylation studies

4.3.1. Effect of sorbent dosage

Figure 5 showed that the sorbent (DRPs) exhibited an initial sharp decrease in WPG with increase in sorbent dose. This could

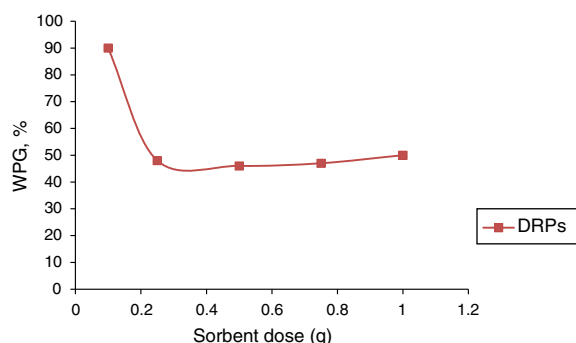


Fig. 5. Effect of sorbent dosage on WPG due to acetylation of DRPs.

be due to the insufficient amount of acetyl groups (present in the constant quantity of acetic anhydride) available for attachment at the –OH sites in the cellulose. As the sorbent dose increases and same amount of acetyl groups are available to be attached to the increasing –OH sites, the WPG due to acetylation will decrease due to the ratio of acetyl groups to –OH decreases. This suggests that the quantity of acetic anhydride needs to be increased as sorbent dose increases, in order to provide large or sufficient acetyl groups for attachment or substitution at the increased –OH sites.

4.3.2. Effect of catalyst

It was observed in Figure 6 that the WPG due to acetylation of the sorbent (DRPs) showed an initial increase, as the catalyst was introduced. It can be seen that WPG due to acetylation of DRPs increased rapidly through a maximum. Thus, the amount of catalyst required for optimum acetylation of DRPs is 1% of the quantity of acetic anhydride used. This is consistent with report by Adebajo and Frost (2004) which states that extent of acetylation for all samples is generally lower in the absence of the catalyst. The sharp increase could possibly be because NBS

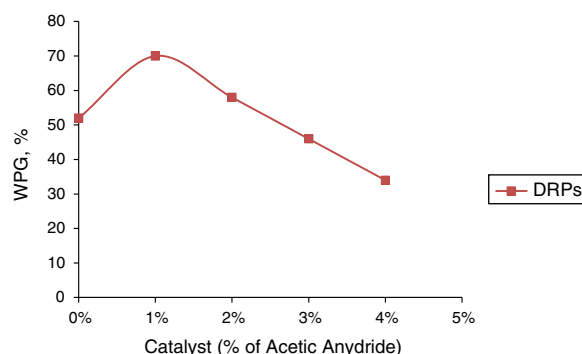


Fig. 6. Effect of catalyst on WPG due to acetylation of DRPs.

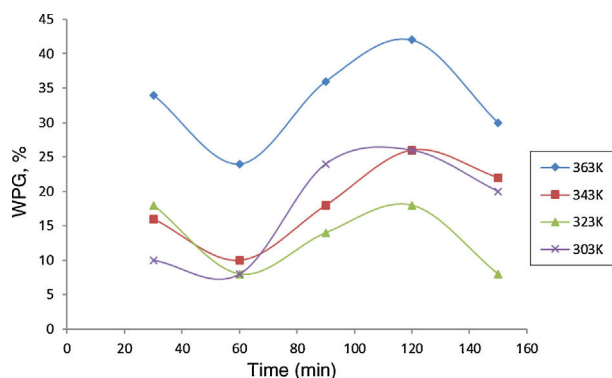


Fig. 7. Effect of reaction time on WPG due to DRPs acetylation.

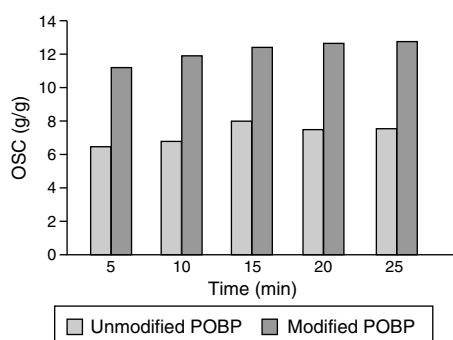


Fig. 8. Oil sorption capacity for unmodified and modified DRPs at various contact time. Key: POBP means *Delonix regia* pods (DRPs).

catalyst increases the rate of reaction by enhancing the electrophilic character of the acyl carbon center (Hill et al., 1998).

4.3.3. Effect of reaction time

The effect of reaction time at various temperatures was studied. Figure 7 represents the effect of reaction time on WPG due to acetylation of DRPs. In acetylation of DRPs, The variation of WPG with reaction time at all the temperatures follows similar pattern with WPG reaching a minimum first at 60 min and then increases to maximum at 120 min before decreasing. After 30 min, the level of acetylation decreased probably due to de-acetylation mechanism as proposed in report on acetylation of commercial cotton (Adebajo & Frost, 2004). Thus, it is possible that the extent of acetylation far exceeded de-acetylation of the DRPs samples beyond 60 min reaction.

4.4. Oil sorption capacity

4.4.1. Effect of contact time

It was observed from Figure 8 that the crude oil sorption capacity of the modified sorbent (DRPs) was higher than that of the unmodified, within the time range studied. This could be attributed to the hydrophobic nature of the sorbent surface after modification, as shown by the FTIR studies. Figure 8 also showed that there was an increase in oil sorption capacity of the sorbents (unmodified and modified), with increase in contact time up to 15 min when the sorption process reached equilibrium. This may be due to adsorption of crude oil on the surface

of the sorbents first, before the crude oil starts to penetrate the inner microscopic voids (Amer, El-Maghraby, Malash, & Nahla, 2007). Similar findings were reported by Hussein, Amer, El-Maghraby, and Taha (2009), Thompson et al. (2010) and Kudaybergenov et al. (2012) on crude oil sorption by cotton fibers, rice husks (acetylated and unacetylated) and thermally treated rice husks respectively.

4.5. One-way factorial analyses

In Figure 7, the variations of weight percent gain (WPG) due to acetylation with reaction time at various temperatures were shown. ANOVA results presented in Table 2 showed that the observed variations of WPG with reaction time have significant difference at different reaction temperatures.

The crude oil sorption capacity values of the unmodified and modified sorbents at various contact time were tested for significant difference to ensure that the variation was not by chance. Table 3 showed that the crude oil sorption capacity of the sorbents were significantly different. Thus, the crude oil sorption capacity of the modified sorbent was significantly higher than that of the unmodified as shown in Figure 8.

It can be concluded from the analysis of variance (ANOVA) results presented in Tables 2 and 3, that at $\alpha = 0.05$, the null hypotheses were rejected because; $F_{\text{cat}} > F_{\text{crit}}$, where $F_{\text{cal}} = 8.71$ (time effect on WPG) and $F_{\text{cal}} = 335.7$ (OSC variation with time).

The tabulated ANOVA results of Tables 2 and 3 infer the following respectively:

- the differences in the analyzed means of WPG due to acetylation of DRPs at reaction times 30–150 min and operating temperatures (303–363 K), were enough to show that statistical/significant difference exist between them; and
- the differences in the analyzed means of crude oil sorption capacity of the unmodified and modified sorbent at contact times 5–25 min, were enough to show that statistical/significant difference exist between them.

4.6. Kinetics

4.6.1. Acetylation kinetics

The kinetics of acetylation was studied by fitting the obtained data from effect of reaction time, in rate curves – pseudo-first order, pseudo-second order, intra-particle diffusion and liquid film-diffusion as shown in Eqs. (3), (5), (7) and (9) respectively. The value of coefficient of determination (R^2) for the different kinetic plots at different temperatures was used to determine which mechanism contributes to acetylation process at each temperature of the kinetic plot. Coefficient of determination (R^2) values within $0.43 \leq R^2 \leq 0.83$ are considered high and moderate (Dowine & Heath, 1974).

Table 4 showed that in acetylation of DRPs, the mechanism involved in acetylation at 323 K and 363 K, cannot be ascertained due to the low values of R^2 . Moderate and high values of R^2 was obtained at 303 K and 343 K which implies that these temperatures are most suitable for acetylating DRPs. The

Table 2

ANOVA result of weight percent gain (WPG) variation with reaction time at various temperature.

Source of variation	SS	df	MS	F_{cal}	F_{crit}	P -value	Inference
Between groups	1136.80	3	378.93	8.71	3.24	0.001	Significant
Within groups	696	16	43.5				
Total	1832.80	19					

Table 3

ANOVA result on variation of crude oil sorption capacity of modified and unmodified DRPs with time.

Source of variation	SS	df	MS	F_{cal}	F_{crit}	P -value	Inference
Between groups	56.83	1	56.83	335.66	5.32	8.11E–08	Significant
Within groups	1.35	8	0.17				
Total	58.19	9					

Table 4

Data from kinetic plots of DRPs acetylation.

Temp (K)	Pseudo-first order			Pseudo-second order			Intra-particle diffusion			Liquid film diffusion		
	k_1	WPG ₀	R^2	k_2	WPG ₀	R^2	k_3	C	R^2	k_4	R^2	Intercept
<i>Delonix regia</i> Pods (DRPs)												
363	–0.0001	29.84	0.051	–0.0007	27.78	0.2000	0.5923	27.76	0.0551	–0.0056	0.1201	1.6430
343	–0.0053	10.83	0.4781	–0.0003	11.63	0.3077	1.5947	3.7541	0.4905^a	0.0044	0.0886	0.4990
323	0.0027	15.78	0.0972	0.0002	14.49	0.0907	–0.618	18.876	0.1076	0.002	0.0228	0.3610
303	–0.0086	7.32	0.5749	–0.0006	7.75	0.5404	2.3273	–3.7736	0.5754^a	0.0053	0.0587	0.5010

^a Means the highest R^2 values. R^2 values that are moderate are bolded.

highest values of R^2 are obtained at intra-particle diffusion plot of 303 K and 343 K. Thus, diffusion into the pores of DRPs is the rate controlling mechanism in the acetylation of DRPs at these temperatures but it is not the sole mechanism because the straight line did not pass through the origin. Table 4 also showed that surface reaction was remarkably involved in DRPs acetylation mechanism at 303 K and 343 K but chemical reaction also contributed to the acetylation mechanism at 303 K. Large intercepts obtained by using intra-particle diffusion model at 303 K and 343 K suggested that the process was largely surface reaction (Eba et al., 2010). This could also be supported by the fact that at these temperatures (303 K and 343 K), the coefficient of determination values for pseudo-first order is higher than that of pseudo-second order kinetics.

High and moderate values of R^2 obtained at room temperature for pseudo first, pseudo second and intra particle diffusion kinetic plots, also suggests that acetylation of DRPs is energy efficient since physical and chemical reaction between the modifier (acetic anhydride and NBS catalyst) and DRPs, can take place at room temperature.

Thus, kinetic studies showed that acetylation of DRPs was temperature dependent and largely by surface reaction.

4.6.2. Crude oil sorption kinetics

In order to elucidate the mechanisms involved in the crude oil sorption, the kinetic results obtained by calculating the oil sorption capacity (OSC) of sorbents, were analyzed by pseudo-first order, pseudo-second order, intra-particle diffusion and liquid

film-diffusion model as proposed in Eqs. (4), (6), (8) and (9) respectively.

Table 5 showed that the coefficients of determination obtained were generally poor for unmodified DRPs. It was observed from Table 5, that the rates of physisorption, chemisorption and transport of solute from liquid to sorbent phase, were faster in the modified DRPs than in the unmodified since pseudo-first order, pseudo-second order and liquid film-diffusion rate constants, were found to be higher in the modified DRPs. This can be attributed to the hydrophobic nature of the modified sorbent (DRPs).

The diffusion rate constant of modified sorbent was observed to be lower than that of unmodified sorbent. This implies that the diffusion of crude oil into the pores of modified sorbent was slower than its diffusion into the pores of unmodified sorbent. This is due to the reduction in the size of the small and large pores of the modified sorbent as a result of acetylation.

The coefficient of determination (R^2) values for crude oil sorption by modified DRPs was found to be moderate for most of the kinetic models. It was observed that intra-particle diffusion was the rate controlling mechanism for the sorption process but it is not the sole mechanism since the line did not pass through the origin.

In the reaction kinetics (i.e. pseudo-first and pseudo-second order), the pseudo-first order rate constants for the sorbents were found to be lower than the pseudo-second order rate constants. This means that surface sorption being the slowest step in the reaction kinetics, is the rate controlling step in the sorption process. This is supported by the fact that in crude oil sorption by modified DRPs, the coefficient of determination values

Table 5

Kinetic data for crude oil sorption by unmodified and modified DRPs.

Sample	Pseudo-first order			Pseudo-second order			Intra-particle diffusion			Liquid film diffusion	
	OSC ₀	R ²	k ₁	OSC ₀	R ²	k ₂	C	R ²	k ₃	R ²	k ₄
Unmodified DRPs	6.76	0.2546	−0.0042	6.84	0.200	−0.0004	6.36	0.2541	0.2260	0.0503	0.0286
Modified DRPs	11.57	0.4932	−0.0024	11.57	0.4717	−0.0002	11.14	0.5837^a	0.2238	0.1014	0.0720

^a Means the highest R² values. R² values that are moderate are bolded.

(which was moderate) for pseudo-first order was higher than that of pseudo-second order. This implies that surface sorption is the rate controlling mechanism in the sorption process and chemisorption was also partly involved in the process. These also show that rate constants are very sensitive to the determination of mechanism of oil sorption process, thus, indicating the reliability of the rate constant technique in predicting oil sorption mechanism.

5. Conclusion

DRPs was successfully modified and this, makes it an effective sorbent for treatment of crude oil contaminated water bodies. Sorbent (DRPs) dosage, catalyst (NBS) concentration, reaction time and temperature played significant roles on the extent of acetylation. Diffusion into pores of the DRPs was largely responsible for its acetylation while the contributing mechanisms depends on the acetylation temperature. DRPs can be acetylated effectively at room temperature (303 K). Surface chemistry of modified DRPs (presence of hydrophobic functional groups) favored its crude oil sorption capacity. Thus, the rate of physical and chemical sorption is faster in the modified DRPs than the unmodified. The scanning electron microscope (SEM) analysis showed that both acetylation and crude oil sorption affect the surface structure, pore sizes and fiber lengths of the sorbent (DRPs). FTIR studies indicates that acetyl functional groups were successfully attached to the sorbent and crude oil was adsorbed at the hydrophobic functional groups.

Conflict of interest

The authors have no conflicts of interest to declare.

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